

DESCRIBING CHEMICAL REACTIONS

Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation

Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical ____1____. The substances that undergo a chemical change are the ____2____. The new substances formed in a chemical reaction are the ____3____. In accordance with the law of conservation of ____4____, a chemical equation must be balanced. When balancing an equation, you place ____5____ in front of reactants and products so that the same number of atoms of each ____6____ are on each side of the equation. An equation must never be balanced by changing the ____7____ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is ____8____; for a solid, ____9____; and for a gas, ____10____. A substance dissolved in water is designated ____11____. If a ____12____ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- _____ 14. The symbol Δ placed over the arrow in an equation means that heat is supplied to the reaction.
- _____ 15. Atoms are destroyed in a chemical reaction.
- _____ 16. A skeleton equation is not a balanced equation.

Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I
- a. Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.
 - b. Acetylene gas (C_2H_2) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

Section Quiz Describing Chemical Reactions Answers

Richard Grime



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Learning and Collaboration Technologies Panayiotis Zaphiris,Andri Ioannou,2015-07-18 The LNCS volume 9192 constitutes the refereed proceedings of the Second International Conference on Learning and Collaboration Technologies LCT 2015 held as part of the 17th International Conference on Human Computer Interaction HCII 2015 in Los Angeles CA USA in August 2015 jointly with 15 other thematically similar conferences The total of 1462 papers and 246 posters presented at the HCII 2015 conferences were carefully reviewed and selected from 4843 submissions These papers address addressing the following major topics technology enhanced learning adaptive and personalised learning and assessment virtual worlds and virtual agents for learning collaboration and Learning Serious Games and ICT in education

USAF Instructors Journal ,1967

Elements of Chemical Reaction Engineering H. Scott Fogler,2020-08-18 The Definitive Guide to Chemical Reaction Engineering Problem Solving With Updated Content and More Active Learning For decades H Scott Fogler s Elements of Chemical Reaction Engineering has been the world s dominant

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