Name: \_\_\_\_\_\_ Date: \_\_\_\_\_\_

# **Honors Chemistry Net Ionic Equations**

Write down the following molecular equations. Balance them and show the ionic and net ionic forms of the equations.

1) strontium bromide(aq.) + strontium sulfate(s) + potassium sulfate(aq.)

Molecular equation:

lonic equation:

Net ionic equation:

2) manganese(II) chloride(aq.) + \_\_\_ manganese(II) carbonate(s) + ammonium carbonate(aq.)

Molecular:

tonic:

Net ionic:

3) cobalt(III) bromide(aq.) + cobalt(III) sulfide(s) + potassium bromide(aq.)

Molecular:

lonic:

Net ionic:

4) barium nitrate(aq.) + \_\_\_\_ barium phosphate(s) + ammonium phosphate(aq.)

Molecular:

Ionic:

Net ionic:

5) lithium hydroxide(aq.) + \_\_\_ lithium chloride (aq.) + barium chloride(aq.)

Molecular:

Ionic:

Net ionic:

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**RM Cervero** 

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