The equilibrium constant for the reaction.



$$H_{2(g)} + I_{2(g)} \rightleftharpoons 2HI_{(g)}$$

at 448 C is 50.

a) If 1.0 mol of H2 is mixed with 1.0 mol of I2 in a 0.50 L container and allowed to react at 448°C, what is the equilibrium [HI]? Inital [H3] = 1.0 mol = 2.0 M, [I] = 2.0M, [HI] = 0

			Not be surrium feature
	H2+	I2 =	ZHI
	2.0	2.0	0
[C]	-×	-×	+ 2×
TE]	2.0-x	2.0-×	2× .
EJ	2.9-1.551	2.0-1559	2(1.559)
TF7	0.44	0.44	3.1

$$Keq = \frac{EHIJ}{EH_2JLI_2J}$$

50. = $\frac{(2x)^2}{(2.0-x)^2}$

Keq = $\frac{\text{EHI}_{3}^{2}}{\text{EH}_{2}\text{I}\text{EI}_{2}\text{I}}$ | 7.071 (2-x) = 2x 14.142-7.071x = 2x 14.142 = 9.071 x x = $\frac{(2x)^{2}}{(2.0-x)^{2}}$ | $x = \frac{14.142}{9.071} = 1.559$

b) How many moles of HI are formed at equilibrium? (Actual yield)

Answer 1.6 moles of HI

Given Keq for the reaction:

is 0.042 at 250°C, what will happen if 2.50 mol of PCIs, 0.600 mol of CI2 and 0.600 mol of PCI₃ are placed in a 1.00 flask at 250°C ? (Will the reaction shift left, right, or not occur at all?)

$$Keq = \frac{EPCI_3 3 ECI_2 3}{EPCI_5 3}$$

Trial $Keq = \frac{(0.600)(0.600)}{(2.50)}$

Trial $Keq = 0.144$

Actual $Keq = 0.042$

- since Trial Keg > Keg The reaction will shift to the LEFT as it approaches equilm So PCIS will increase while EPCh3 and ECh3 decrease.

Answer LHI] = 3.1 M

Answer Shift LEFT

Nelson Chemistry Exercise Answers

Judd E. Hollander

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