

Chemistry 11

Chapter 4 – The MOLE

Atomic Masses and Avogadro's Hypothesis

Experimental work by the English chemist **John Dalton** (1766-1844) was concerned with how much of one element could combine with a given amount of another element. He suggested the following hypotheses:

- Molecules are made up of "atoms" of various elements
- If compound B contains twice the mass of element X as does compound A, then compound B must contain twice as many atoms as X
- Simple compounds are made up of only one atom of the two elements making up the compound (Dalton was aware that this hypothesis might be quite wrong but he had no way of knowing how many atoms were actually contained in a compound)

Dalton did not attempt to figure out the mass of an individual atom of any element. Instead he simply assigned an **arbitrary mass** to each element, assuming that hydrogen was the lightest element and therefore could be assigned a mass of "1". Since his experiments found that C was 6 times heavier than H, C was assigned a mass of 6. Similarly, O was 16 times heavier than H and was assigned a mass of "16". In this way, Dalton was able to calculate "relative masses" for several different elements.

Example: The reaction between 2.74 g of hydrogen gas and 97.26 g of chlorine gas makes 100 g of hydrogen chloride. Assuming hydrogen chloride contains one atom each of hydrogen and chlorine, the chlorine atom is

Exercise:

1. Assume a chemist working in 1810 (Mr. Pong – you're *THAT* old?) tries to duplicate some of Dalton's results, published a year or two earlier, and obtains the following results (expressed in modern mass unit, grams)

11.1 g of hydrogen gas react with 88.9 g of oxygen gas
46.7 g of nitrogen gas react with 53.3 g of oxygen gas
42.9 g of carbon react with 57.1 g of oxygen gas

Assuming a mass of "1" for hydrogen, calculate the relative masses of oxygen, nitrogen and carbon. (Don't be surprised if the values you calculate are not what you would expect from the periodic table. Not all the molecules produced in reactions involve 1:1 combinations of atoms, which was a problem plaguing early chemists for many years.

Dalton's atomic mass scale was partly in error because not all of the molecules he studied actually contained only one atom of each element. For example, it was now known that a molecule of hydrogen contains two atoms. Dalton's mass scale was just being introduced when the French chemist **Joseph Gay-Lussac** began to study how gases reacted.

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Tompkins, Dick, Di Giuseppe, Maurice



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The British National Bibliography, 1965 The British National Bibliography Arthur James Wells, 1970

The United States Catalog Mary Burnham, Carol Hurd, 1928 *The United States Catalog*, 1928 Nelson Chemistry, Atlantic Edition. Teacher's Resource (chapter 1-9) Frank Jenkins, 1996 **Nelson Chemistry 11** Tompkins, Dick, Di Giuseppe, Maurice, 2002 *Nelson Chemistry* Frank Jenkins, 1996

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Table of Contents Nelson Chemistry 1answers Chapter 6

1. Understanding the eBook Nelson Chemistry 1answers Chapter 6
 - The Rise of Digital Reading Nelson Chemistry 1answers Chapter 6
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 1answers Chapter 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 1answers Chapter 6
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 1answers Chapter 6
 - Personalized Recommendations
 - Nelson Chemistry 1answers Chapter 6 User Reviews and Ratings
 - Nelson Chemistry 1answers Chapter 6 and Bestseller Lists
5. Accessing Nelson Chemistry 1answers Chapter 6 Free and Paid eBooks
 - Nelson Chemistry 1answers Chapter 6 Public Domain eBooks
 - Nelson Chemistry 1answers Chapter 6 eBook Subscription Services
 - Nelson Chemistry 1answers Chapter 6 Budget-Friendly Options
6. Navigating Nelson Chemistry 1answers Chapter 6 eBook Formats

- ePub, PDF, MOBI, and More
- Nelson Chemistry 1answers Chapter 6 Compatibility with Devices
- Nelson Chemistry 1answers Chapter 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Chemistry 1answers Chapter 6
 - Highlighting and Note-Taking Nelson Chemistry 1answers Chapter 6
 - Interactive Elements Nelson Chemistry 1answers Chapter 6
- 8. Staying Engaged with Nelson Chemistry 1answers Chapter 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 1answers Chapter 6
- 9. Balancing eBooks and Physical Books Nelson Chemistry 1answers Chapter 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 1answers Chapter 6
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nelson Chemistry 1answers Chapter 6
 - Setting Reading Goals Nelson Chemistry 1answers Chapter 6
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nelson Chemistry 1answers Chapter 6
 - Fact-Checking eBook Content of Nelson Chemistry 1answers Chapter 6
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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