

Stuff Stacking Up? Donate to The Salvation Army!

Call for a FREE Pick-Up!
1-800-SA-TRUCK

Don't let
that **STUFF**
go to
WASTE!

Join The Salvation Army in our campaign - **DOING THE MOST GOOD!** Donate gently used clothing, shoes, appliances, household goods, electronics, home and office furniture. Your donation is always tax deductible and provides much needed product to our 9 Orange County Family Thrift Stores helping to fund vital programs in your community.

In the last year, we have diverted over
2 million lbs from Orange County landfills!

3 Easy Steps:

- 1** **DESIGNATE** which items you would like to donate
- 2** **CALL US** at 1-800-SA-TRUCK to schedule a Pick-Up anytime Monday - Friday
- 3** **COLLECT** your receipt for Tax-deduction from your Driver

Then, come **SHOP** with us to help us provide our no-cost services! Our Family Stores are known for quality, affordability and great deals on home decor, furniture, clothing and vintage or hard to find items. We have 9 convenient Orange County stores to shop.

Want to Host Your Own Drive? Contact Us:

The Salvation Army, Anaheim ARC

1300 South Lewis Street

(714) 758-0414 ext. 3438

Community Relations Manager – Danny Hauger

Interested in a companywide Clothing Drive or Shoe Drive Fundraiser? We can help with that too!



**SHOP
WITH US**

ANAHEIM

1100 N. La Palma Pkwy.
(714) 535-4631

BUENA PARK

7035 Stanton Blvd.
(714) 521-7390

COSTA MESA

2126 Harbor Blvd.
(949) 642-3636

FOUNTAIN VALLEY

16119 Brookhurst St.
(714) 775-7124

HUNTINGTON BEACH

17362 Beach Blvd.
(714) 841-2896

LAKE FOREST

23820 Mercury Rd.
(949) 581-1605

ORANGE

180 S. Tustin Blvd.
(714) 532-2235

SAN CLEMENTE

2727 Via Cascadita
(949) 492-0133

SANTA ANA

2601 West First St.
(714) 541-6726

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Salvation Army 2013 Donation Guide

M Woodhall



Salvation Army 2013 Donation Guide:

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sep 24 2023 one footed tricks performing tricks with only one foot strapped into the board hand drag dragging one or both hands on the snow while performing a trick or turn miller flip basically an inverted 360 with a front hand hand drag

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trick 1 ollie an ollie is a specific way of jumping off of the ground vertically while snowboarding in general we recommend learning how to ollie first as it is a critical step towards learning other snowboard tricks once the ollie is mastered it can be used for flatland tricks rail tricks and jump tricks

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mar 14 2023 1 don t attempt tricks on day one 2 bend your knees 3 ride across slope not downhill 4 always watch out for other people 5 start small 6 wear warm snow approved clothing 7 make sure

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jul 12 2021 1 ollie the ollie is one of the most basic and most essential tricks for every snowboarder to learn it involves jumping on the snow with your board strapped on and is a foundational part of many more complex maneuvers the ollie gets its name from skateboarding where the trick originated and is pretty much the same thing

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oct 18 2023 1 fitness first snowboarding is excellent exercise however if you try to go directly from the couch to the slopes you will be exhausted well before your first lesson is done you do not have to be a triathlete but you should be a bit active be it working out at the gym swimming laps or taking regular hikes

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dec 27 2022 there are a variety of snowboard tricks that can be performed on a snowboard and each trick has its own unique name some of the most popular snowboard tricks include the frontside air this trick is performed by jumping off of the front side of the snowboard and performing a 360 degree rotation in the air before landing back on the snowboard

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some of the most popular spin tricks include an alley oop half cab and hard way other snowboarding tricks the above list is just a glimpse into the many styles and types of tricks that you can perform on a snowboard other popular tricks that you ll

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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web hess s law labs by austin lee alayna baron lily zmachinski introduction in order to calculate the enthalpy change for the combustion of magnesium oxide $\text{mg s} + \frac{1}{2}\text{O}_2 \text{g} \rightarrow \text{MgO s}$ we used

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web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows $\text{NaOH s} + \text{Na aq} \rightarrow \text{OH aq}$

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web hess s laws abstract the following experiment was designed using hess s law and simple calorimetry to determine the enthalpy for NaOH s hess s law says that the overall change in enthalpy for the reaction is the sum of all changes regardless of the different stages or phases of a reaction

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web hess s law states that if a reaction is the sum of two or more other reactions the Δh for the overall process must be the sum of the Δh values of the constituent reactions enthalpy change Δh is independent of the path that a reaction follows to move from reactants

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web this activity provides a demonstration of hess law using three reactions the solubility NaOH in water the solubility NaOH in HCl and the reaction of a solution of HCl and a solution of NaOH

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web hess s law worksheet answers 1 calculate h for the reaction $\text{C}_2\text{H}_4 \text{g} + \text{H}_2 \text{g} \rightarrow \text{C}_2\text{H}_6 \text{g}$ from the following data $\text{C}_2\text{H}_4 \text{g} + 3\text{O}_2 \text{g} \rightarrow 2\text{CO}_2 \text{g} + 2\text{H}_2\text{O l}$ h 1411 kJ $\text{C}_2\text{H}_6 \text{g} + 3\frac{1}{2}\text{O}_2 \text{g} \rightarrow 2\text{CO}_2 \text{g} + 3\text{H}_2\text{O l}$ h 1560 kJ

7a first law enthalpy calorimetry and hess s law worksheet - Oct 05 2022

web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction c 2 h 4

experiment 9 enthalpy of reaction hess s law - Jul 14 2023

web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} + \text{NaOH (aq)} \rightarrow \text{H}_2\text{O (l)}$ 2 $\text{NaOH (s)} + \text{HCl (aq)} \rightarrow \text{H}_2\text{O (l)}$ 3 $\text{NaOH (aq)} + \text{HCl (aq)} \rightarrow \text{H}_2\text{O (l)}$