

## Calculus-Based Properties (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose,  $C_6H_{12}O_6$ , and 100 g of water? Equations:  $K_f = 1.86^\circ C/m$ ,  $K_b = 0.512^\circ C/m$ .

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 5.56 \text{ mol/kg} = 10.3^\circ C \text{ (freezing)}$$

$\Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 5.56 \text{ mol/kg} = 2.85^\circ C$   
 Boiling point =  $100^\circ C + 2.85^\circ C = 102.85^\circ C$   
 Freezing point =  $0^\circ C - 10.3^\circ C = -10.3^\circ C$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea and 100 g of water? Equations:  $K_f = 1.86^\circ C/m$ ,  $K_b = 0.512^\circ C/m$ .  $K_f$  for urea =  $1.86^\circ C/m$ ,  $K_b$  for urea =  $0.512^\circ C/m$ .

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60 \text{ g}}}{0.100 \text{ kg}} = 16.7 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 16.7 \text{ mol/kg} = 31.1^\circ C$$

$\Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 16.7 \text{ mol/kg} = 8.55^\circ C$   
 Boiling point =  $100^\circ C + 8.55^\circ C = 108.55^\circ C$   
 Freezing point =  $0^\circ C - 31.1^\circ C = -31.1^\circ C$

3. How many grams of  $C_6H_{12}O_6$  must be added to 100 g of  $H_2O$  to raise the boiling-point of the solution to  $102.5^\circ C$ ? Assume that the solute does not dissociate. Equations:  $K_b = 0.512^\circ C/m$

$$\Delta T_b = K_b \cdot m \Rightarrow 2.5^\circ C = 0.512^\circ C/m \cdot m \Rightarrow m = 4.88 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} \Rightarrow 4.88 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$\Rightarrow \text{moles solute} = 0.488 \text{ mol}$$

$$\text{Mass} = 0.488 \text{ mol} \times 180 \text{ g/mol}$$

$$\text{Mass} = 87.8 \text{ g}$$

$$\text{Total mass} = 100 \text{ g} + 87.8 \text{ g} = 187.8 \text{ g}$$

$$\text{Mass} = 87.8 \text{ g}$$

$$\text{Total mass} = 100 \text{ g} + 87.8 \text{ g} = 187.8 \text{ g}$$

4. A solution composed from 10 g of glucose,  $C_6H_{12}O_6$ , and 100 g of water is prepared. The freezing-point of this solution is  $-1.7^\circ C$ . The normal freezing point of pure water is  $0^\circ C$  and the freezing-point depression constant for water is  $1.86^\circ C/m$ . Determine the molar mass of the compound. Assume it is a nonelectrolyte. Equations:  $K_f = 1.86^\circ C/m$

$$\Delta T_f = K_f \cdot m \Rightarrow 1.7^\circ C = 1.86^\circ C/m \cdot m \Rightarrow m = 0.914 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} \Rightarrow 0.914 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$\Rightarrow \text{moles solute} = 0.0914 \text{ mol}$$

$$\text{Mass} = 0.0914 \text{ mol} \times 180 \text{ g/mol}$$

$$\text{Mass} = 16.5 \text{ g}$$

$$\text{Total mass} = 100 \text{ g} + 16.5 \text{ g} = 116.5 \text{ g}$$

$$\text{Mass} = 16.5 \text{ g}$$

$$\text{Molar mass} = \frac{\text{mass}}{\text{moles}} = \frac{16.5 \text{ g}}{0.0914 \text{ mol}} = 180 \text{ g/mol}$$

$$\text{Molar mass} = 180 \text{ g/mol}$$

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