Gas Laws Worksheet

atm = 760.0 mm Hg = 101.3 kPa= 760.0 torr

Boyle's Law Problems:

PV. TRV

 If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(749-4)(22.51) = (725-4) V_2$$

 $V_2 = \frac{(749-4)(22.51)}{(725-4)}$

V2 - 23.2L

A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L.
 What is the pressure in the container if the temperature remains constant?

 What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

 A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

Practice With Gas Law Calculations Iii

CO Houle

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