



NEW JERSEY CENTER FOR TEACHING & LEARNING

Aqueous Equilibria: Part 1 - Buffers & Titrations

PSI AP Chemistry

Name _____

Multiple Choice questions

Common-Ion Effect

1) The pH of a solution that contains 0.818 M acetic acid ($K_a = 1.77 \times 10^{-5}$) and 0.172 M sodium acetate is _____.

- A) 4.077 B) 5.434 C) 6.571 D) 8.370 E) 9.922

2) Calculate the pH of a solution prepared by dissolving 0.37 mol of formic acid (HCO_2H) and 0.23 mol of sodium formate (NaCO_2H) in water sufficient to yield 1.00 L of solution. The K_a of formic acid is 1.77×10^{-4} .

- A) 2.1 B) 10.46 C) 3.55 D) 2.31 E) 3.95

3) Calculate the pH of a solution prepared by dissolving 0.25 mol of benzoic acid ($\text{C}_6\text{H}_5\text{O}_2\text{H}$) and 0.15 of sodium benzoate ($\text{NaC}_6\text{H}_5\text{O}_2$) in water sufficient to yield 1.00 L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

- A) 4.41 B) 2.40 C) 3.97 D) 10.08 E) 4.19

4) Calculate the pH of a solution prepared by dissolving 0.15 mol of benzoic acid (HBz) and 0.30 mol of sodium benzoate in water sufficient to yield 1.00L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

- A) 2.52 B) 3.89 C) 4.49 D) 10.16 E) 4.20

5) Calculate the pH of a solution prepared by dissolving 0.750 mol of NH_3 and 0.250 mol of NH_4Cl in water sufficient to yield 1.00 L of solution. The K_b of ammonia is 1.77×10^{-5} .

- A) 5.22 B) 4.27 C) 9.73 D) 8.78 E) 0.89

6) The pH of a solution prepared by dissolving 0.350 mol of solid methylamine hydrochloride ($\text{CH}_3\text{NH}_3\text{Cl}$) in 1.00L of 1.10M methylamine (CH_3NH_2) is _____. The K_b for methylamine is 4.4×10^{-4} . (Assume the final volume is 1.00 L.)

- A) 1.86 B) 2.66 C) 10.15 D) 11.14 E) 10.64

Buffers

7) What does a buffer do?

- a) Keeps the pH of a solution constant.

Practice Multiple Choice For Chem Aqueous Equilibria

Lingjun Ying



Practice Multiple Choice For Chem Aqueous Equilibria:

Practice Multiple Choice For Chem Aqueous Equilibria Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Practice Multiple Choice For Chem Aqueous Equilibria**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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