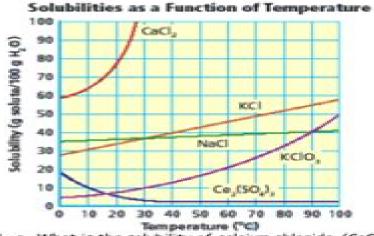
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Solubility Curve Worksheet 2



<u>Directions</u>: Use the graph below to answer the following questions.



- Points that collectively make up the solubility curve (points ON the curve) represent ______ solutions.
- > Points that are BELOW the curve represent solutions.
- > Points ABOVE the solubility curve represent
 solutions, and the difference
 between the point above the curve and on the curve
 represents the amount of solute which will precipitate out.
- Say if the resulting solutions would be saturated, supersaturated or unsaturated.
 - 60 g of KCl at 70°C
 - 10 g of KClO₂ at 60 °C
 - o 70 g of CaCl₂ at 20 °C
- 1. a. What is the solubility of calcium chloride (CaCl.) at 5°C?
 - b. What is the solubility of calcium chloride at 25°C?
 - c. What is happening to the solubility and temperature for CaCl₂? Solubility is ______ as temperature

is _____

- a. What is the solubility of <u>cerium sulfate Ce₂(SO₄)</u>, at 10°C?
 - b. What is the solubility of cerium sulfate at 50°C?
 - c. What is happening to the solubility and temperature for Ce₂(SO₄)₃? Solubility is ______ as

temperature is ______

- Which substance on the graph is least soluble at 10°C?
- 4. Which substance on the graph shows the least change in solubility from 0°C to 100°C?
- 5. Which solution is more concentrated:
 - a) At 10°C, a saturated solution of NaCl or a saturated solution of CaCl₂ (circle one)
 - b) At 55°C, a saturated solution of KClO3 or a saturated solution of Ce2(SO4)3 (circle one)
- 6. What is the mass of cerium sulfate that will dissolve in 50 a of water at 10 degrees Celsius?
- 7. What is the mass of potassium chloride that will dissolve in 250 g of water at 70 degrees Celsius?
- 8. At 90 degrees Celsius, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? How do you know?
- 9. A saturated solution of potassium chlorate is dissolved in 100g of water. If the saturated solution is cooled from 90 degrees Celsius to 60 degrees Celsius, how many grams would crystallize out?

Practice Solubility Problems

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