

Lesson- Metamorphic Rocks

- CORNELL NOTES
- 2 WARM-UPS

- ACTIVITY
- EVALUATION SLIDE
- PRESENTATION

Warm-up

Remember For each classroom slide, students is a 30-second timer. Students will only use this time to engage the understanding.

Slide 1: Warm-up: What is a metamorphic rock? (10 minutes)

Slide 2: Warm-up: What is a metamorphic rock? (10 minutes)

Slide 3: Warm-up: What is a metamorphic rock? (10 minutes)

Activity: Managing Metamorphic Rocks

Students will be able to identify the different types of metamorphic rocks and their characteristics. They will also be able to explain the process of metamorphism and how it affects the original rock.

Students will be able to identify the different types of metamorphic rocks and their characteristics. They will also be able to explain the process of metamorphism and how it affects the original rock.



Metamorphic Rocks

Metamorphic rocks are formed from pre-existing rocks (igneous, sedimentary, or metamorphic) through the process of metamorphism. This process involves changes in temperature, pressure, and chemical environment, which alter the mineral composition and texture of the original rock.

Metamorphic rocks are classified based on their parent rock (protolith) and the conditions of metamorphism. The main types of metamorphic rocks are:

- Sedimentary Metamorphic Rocks:** These rocks are formed from sedimentary rocks. Examples include slate (from shale), schist (from sandstone), and gneiss (from granite).
- Igneous Metamorphic Rocks:** These rocks are formed from igneous rocks. Examples include marble (from limestone) and quartzite (from quartzite).
- Metamorphic Rocks:** These rocks are formed from metamorphic rocks. Examples include eclogite (from eclogite) and amphibolite (from amphibolite).



Lesson Evaluation Sheet

Question	Answer	Score
1. What is a metamorphic rock?		
2. How is a metamorphic rock formed?		
3. What are the three main types of metamorphic rocks?		
4. What is the difference between a sedimentary metamorphic rock and an igneous metamorphic rock?		
5. What is the difference between a metamorphic rock and a sedimentary rock?		
6. What is the difference between a metamorphic rock and an igneous rock?		
7. What is the difference between a metamorphic rock and a sedimentary metamorphic rock?		
8. What is the difference between a metamorphic rock and an igneous metamorphic rock?		
9. What is the difference between a metamorphic rock and a sedimentary metamorphic rock?		
10. What is the difference between a metamorphic rock and an igneous metamorphic rock?		



Rocks Guided And Study Answers Metamorphic

**Del Franz,Phyllis Dutwin,Richard
Ku,Kathleen Peno,Courtney Mayer**



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jan 3 2023 a net ionic equation is a chemical equation that shows only the species participating in a chemical reaction and omits the spectator ions spectator ions are ions that

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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jun 23 2022 what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows

$$4\text{Ag}^+_{\text{aq}} + 2\text{F}^-_{\text{aq}} + 2\text{NH}_4^+_{\text{aq}} + \text{Cr}_2\text{O}_7^{2-}_{\text{aq}} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+_{\text{aq}} + 2\text{F}^-_{\text{aq}}$$

because two NH_4^+ are needed to balance the charge

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