

## Quiz & Worksheet - Determining the Chemical Formula of Hydrates from Empirical Data

1. What is the difference between a hydrate and an anhydrate?

- ☐ An anhydrate is a hydrate with the water removed.
- ☐ A hydrate is an anhydrate with the water removed.
- ☐ A hydrate is necessary for producing water molecules, but an anhydrate isn't.
- ☐ An anhydrate is necessary for producing water molecules, but a hydrate isn't.

2. In the formula  $\text{MgSO}_4 \cdot n\text{H}_2\text{O}$ , what does the  $n$  before the  $\text{H}_2\text{O}$  stand for?

- ☐ The number of moles of water.
- ☐ The number of moles of hydrogen.
- ☐ The number of moles of hydrate.
- ☐ The number of moles of anhydrate.

3. If there are three moles of  $\text{H}_2\text{O}$  for every one mole of  $\text{LiClO}_4$ , then what is the hydrate formula?

- ☐  $\text{LiClO}_4 \cdot 3\text{H}_2\text{O}$
- ☐  $3\text{LiClO}_4 \cdot 3\text{H}_2\text{O}$
- ☐  $\text{LiClO}_4 \cdot (3\text{H}_2\text{O})_3$
- ☐  $3(\text{LiClO}_4) \cdot 3\text{H}_2\text{O}$

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# Practicing Hydrate Formulas Quizzes

**CH Cherryholmes**

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**CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09. The book itself contains chapter-length subject reviews on every subject tested on the AP Chemistry exam, as well as both sample multiple-choice and free-response questions at each chapter's end. Two full-length practice tests with detailed answer explanations are included in the book.

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