

Specific Heat

DIRECTIONS: Use formula $q = (m)(\Delta T)(C)$ to solve the following problems. Show all work and units. Use <http://weebly.com> for specific heat of various substances.

1. A 30g piece of metal absorbs 1,200 joules of heat energy and its temperature change from 25°C to 175°C. Calculate the specific capacity of the metal. What is the likely metal?
2. How many joules of heat are needed to raise the temperature of 15.0 g of iron from 11°C to 68°C, if the specific heat of iron is 0.45 J/g°C?
3. To what temperature will a 50.0 g piece of brass raise if it absorbs 6 kilojoules of heat and its specific heat capacity is 0.38 J/g°C? The initial temperature of the brass is 20.0°C.
4. Identify the substance if 1500.0 g of the substance absorbs 74,500 joules of heat, and its temperature changes from 25°C to 47°C.
5. 100.0 g of 1.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.
6. 25.0 g of sodium is heated from 30°C to 160°C, and absorbs 750 joules of heat in the process. Calculate the specific heat capacity of sodium.
7. What is the specific heat capacity of quartz glass if 200.0 g of the glass absorbs 2520 joules of heat and the temperature rises 18.0°C?
8. If a sample of titanium is initially at 25°C, what is its final temperature if 250.0 g of titanium absorbs 1.0 kilojoules of heat, and the specific heat of titanium is 0.523 J/g°C?
9. How much energy must be absorbed by 300.0 g of steel to increase its temperature from 100.0°C to 450.0°C? (C of steel = .48 J/g °C)
10. When 55.0 g of platinum drops its temperature from 300.0°C to 250.0°C, how much heat energy is released? (C of platinum = .133 J/g °C)
11. How much energy is required to heat 125.0 g of lithium from 2.0°C to 30.0°C? (C of lithium = 3.6 J/g°C)
12. How much heat (in kJ) is given out when 350.0 g of lead cools from 200.0°C to 30.0°C?

Practice Problems For Specific Heat

R. N. Das Gupta



Practice Problems For Specific Heat:

This book delves into Practice Problems For Specific Heat. Practice Problems For Specific Heat is a vital topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Practice Problems For Specific Heat, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Practice Problems For Specific Heat
 - Chapter 2: Essential Elements of Practice Problems For Specific Heat
 - Chapter 3: Practice Problems For Specific Heat in Everyday Life
 - Chapter 4: Practice Problems For Specific Heat in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, this book will provide an overview of Practice Problems For Specific Heat. This chapter will explore what Practice Problems For Specific Heat is, why Practice Problems For Specific Heat is vital, and how to effectively learn about Practice Problems For Specific Heat.
 3. In chapter 2, the author will delve into the foundational concepts of Practice Problems For Specific Heat. The second chapter will elucidate the essential principles that need to be understood to grasp Practice Problems For Specific Heat in its entirety.
 4. In chapter 3, the author will examine the practical applications of Practice Problems For Specific Heat in daily life. The third chapter will showcase real-world examples of how Practice Problems For Specific Heat can be effectively utilized in everyday scenarios.
 5. In chapter 4, this book will scrutinize the relevance of Practice Problems For Specific Heat in specific contexts. This chapter will explore how Practice Problems For Specific Heat is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Practice Problems For Specific Heat. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Practice Problems For Specific Heat.

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