

Stoichiometry

Lecture Examples Answer Key

Ex. 1 Balance the following chemical equations:



Ex. 2 Write a balanced chemical reaction for the combustion of octane, $\text{C}_8\text{H}_{18(l)}$, in air.



Ex. 3 Write the balanced chemical equation for the combination of the metallic element calcium with the nonmetallic element oxygen, O_2 .



Ex. 4 Calculate the molecular weight of

(a) N_2O_5 **MW = 108.0104 g mol⁻¹**

(b) $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ **MW = 158.167 g mol⁻¹**

(c) $[\text{K}_2(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{H}_2\text{O}]$ **MW = 902.1758 g mol⁻¹**

Ex. 5 Calculate the percentage by mass of oxygen in 4a and 4b above.

(a) $[(5 \times 15.9994 \text{ g mol}^{-1}) / 108.0104 \text{ g mol}^{-1}] \times 100\% = \mathbf{74.0642\%}$

(b) $[(4 \times 15.9994 \text{ g mol}^{-1}) / 158.167 \text{ g mol}^{-1}] \times 100\% = \mathbf{40.4620\%}$

Ex. 6 (a) Calculate the number of O atoms in 0.470 mol of $\text{C}_6\text{H}_{12}\text{O}_6$.

$$\begin{aligned} \# \text{ O atoms} &= (0.470 \text{ mol C}_6\text{H}_{12}\text{O}_6)(6.022 \times 10^{23} \text{ molecules/1 mol})(6 \text{ O atoms / 1 molecule}) \\ &= \mathbf{1.70 \times 10^{24} \text{ O atoms}} \end{aligned}$$

(b) Calculate the total number of ions in 38.1 g of CaF_2 .

$$\begin{aligned} \# \text{ ions} &= (38.1 \text{ g CaF}_2)(1 \text{ mol / 78.075 g})(6.022 \times 10^{23} \text{ molecules/1 mol})(3 \text{ ions / 1 molecule}) \\ &= \mathbf{8.82 \times 10^{23} \text{ ions}} \end{aligned}$$

Ex. 7 (a) What is the mass in grams of 1 mol of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

$$\text{mass} = (1 \text{ mol})(180.1572 \text{ g/1 mol}) = \mathbf{180.1572 \text{ g}}$$

(b) What is the mass in grams of 3.52 mol of chromium(III) sulfate decahydrate?

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