

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.*

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.*

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.*

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O   C   **Be**   **Ne**   Same Period
  - Na   Rb   **Fr**   **H**   Same Group
  - Pb**   **C**   Sn   Si   Same Group
  - Au   W   S   **Fr**   **Ne**   Zn   Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O   C   **Be**   **Ne**   Same Period
- Na   Rb   **Fr**   **H**   Same Group
- Pb**   **C**   Sn   Si   Same Group
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# Periodic Trends Practice Answers

**Morris Hein, Susan Arena**



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