Period Date Tuesday, November 1 2016 Name

Periodic Trends

ATOMIC RADIUS

What trend in atomic radius do you see as you go down a group/family on the periodic table?

2. What causes this trend?

3. What trend in atomic radius do you see as you go across a period/row on the periodic table?

4. What causes this trend?

Circle the atom in each pair that has the largest atomic radius.

a) Al B

b) S O c) Br Cl

d) Na Al e) O F f) Mg Ca

Put the following elements in order from smallest to largest atomic radius and explain why: C. O. Sn. Sr.

ELECTRONEGATIVITY

Define electronegativity

8. How does the ionic radius of a nonmetal compare with its atomic radius?

9. What trend in electronegativity do you see as you go down a group/family on the periodic table?

10. What causes this trend?

11. What trend in electronegativity do you see as you go across a period/row on the periodic table?

12. What causes this trend?

Circle the atom in each pair that has the greater electronegativity.

a) Ca Ga b) Li O c) Cl S d) Br As

e) Ba Sr

f) O S

GENERAL QUESTIONS

14. Which group tends to form +1 ions?

15. Which group tends to form +2 ions?

16. Which group tends to form -1 ions?

17. Which group tends not to form ions or react? _____

Periodic Trends Answer Key

Rachel Sandford

Periodic Trends Answer Key:

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Table of Contents Periodic Trends Answer Key

- 1. Understanding the eBook Periodic Trends Answer Key
 - The Rise of Digital Reading Periodic Trends Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Answer Key
 - Personalized Recommendations
 - Periodic Trends Answer Key User Reviews and Ratings
 - Periodic Trends Answer Key and Bestseller Lists
- 5. Accessing Periodic Trends Answer Key Free and Paid eBooks

- Periodic Trends Answer Key Public Domain eBooks
- Periodic Trends Answer Key eBook Subscription Services
- Periodic Trends Answer Key Budget-Friendly Options
- 6. Navigating Periodic Trends Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Periodic Trends Answer Key Compatibility with Devices
 - Periodic Trends Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Answer Key
 - Highlighting and Note-Taking Periodic Trends Answer Key
 - Interactive Elements Periodic Trends Answer Key
- 8. Staying Engaged with Periodic Trends Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Answer Key
- 9. Balancing eBooks and Physical Books Periodic Trends Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Answer Key
 - Setting Reading Goals Periodic Trends Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Answer Key
 - Fact-Checking eBook Content of Periodic Trends Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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