

Study Guide: Periodic Table and Periodic Trends (Honors)

1. Differentiate between Mendeleev's periodic table and Moseley's periodic table.

Mendeleev put the elements in order by atomic mass, but he noticed that in several places, the atomic masses did not always increase sequentially. Moseley discovered that ordering by atomic number was the correct way to describe the order.

2. State the periodic law.

The properties of the elements are periodic functions of their atomic numbers.

3. Vertical columns on the periodic table are called **groups or families** and horizontal rows are called **periods**.

4. Elements with similar chemical properties are placed in the same **group**.

5. Complete the table below.

Electron Config.	Block	Period	Group/Type	Element
[Xe] 6s ² 4f ¹⁴ 5d ¹⁰ 6p ⁶	p	6	18 / Noble Gas	Radon
[Xe] 6s ¹	s	6	1 / Alkali Metal	Cesium
[Ar] 4s ¹ 3d ⁵	d	4	6 / Trans. Metal	Chromium
[He] 2s ²	s	2	2 / Alk. Earth Metal	Beryllium
[Kr] 5s ¹ 4d ¹⁰	d	5	11 / Trans. Metal	Silver
[He] 2s ² 2p ⁵	p	2	17 / Halogen	Fluorine

6. Determine whether the following statements are describing alkali metals (A), alkaline earth metals (AE), transition metals (TM), halogens (H), noble gases (NG):

- a) most reactive non-metals **H**
- b) s²p⁶ valence configuration **NG**
- c) second most reactive metals **AE**
- d) s¹ valence configuration **A**
- e) exceptions found to e- configurations **TM**
- f) soft metals **A**
- g) d orbitals being filled **TM**

7. Would you expect Li and S to have different chemical properties? Why? **Yes; Li is an alkali metal with 1 valence e-, and S is a nonmetal with 6 valence e-.**

8. Explain why a magnesium atom is smaller than: a sodium atom; a calcium atom. **Mg is smaller than Na because the greater nuclear charge in Mg pulls the e- cloud in tighter (periodic trend). Mg is smaller than Ca because it has fewer energy levels (group trend).**

9. Predict the size of the astatine atom compared to the sulfur atom. **At is larger than S, because At is 3 periods down from S, but only 1 group over. The trends compete, but the sizeable increase in energy levels overcomes the slight difference in nuclear charge.**

10. Would you expect a Cl⁻ ion to be larger or smaller than an Mg²⁺ ion. Explain. **A Cl⁻ ion is larger than a Mg²⁺ ion, because the chloride ion has 3 energy levels, matching**

Periodic Trends Study Guide

G Orfield



Periodic Trends Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

ACT Study Guide Premium Prep, 2024: 6 Practice Tests + Comprehensive Review + Online Practice Barron's Educational Series, Brian Stewart, 2024-01-02 Expert tips and study advice to prepare you for the ACT test

Organic Chemistry Study Guide Robert J. Ouellette, J. David Rawn, 2014-11-04 Organic Chemistry Study Guide Key Concepts Problems and Solutions features hundreds of problems from the companion book Organic Chemistry and includes solutions for every problem Key concept summaries reinforce critical material from the primary book and enhance mastery of this complex subject Organic chemistry is a constantly evolving field that has great relevance for all scientists not just chemists For chemical engineers understanding the properties of organic molecules and how reactions occur is critically important to understanding the processes in an industrial plant For biologists and health professionals it is essential because nearly all of biochemistry springs from organic chemistry Additionally all scientists can benefit from improved critical thinking and problem solving skills that are developed from the study of organic chemistry Organic chemistry like any skill is best learned by doing It is difficult to learn by rote memorization and true understanding comes only from concentrated reading and working as many

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Table of Contents Periodic Trends Study Guide

1. Understanding the eBook Periodic Trends Study Guide
 - The Rise of Digital Reading Periodic Trends Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Study Guide
 - Personalized Recommendations
 - Periodic Trends Study Guide User Reviews and Ratings

- Periodic Trends Study Guide and Bestseller Lists
- 5. Accessing Periodic Trends Study Guide Free and Paid eBooks
 - Periodic Trends Study Guide Public Domain eBooks
 - Periodic Trends Study Guide eBook Subscription Services
 - Periodic Trends Study Guide Budget-Friendly Options
- 6. Navigating Periodic Trends Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Study Guide Compatibility with Devices
 - Periodic Trends Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Study Guide
 - Highlighting and Note-Taking Periodic Trends Study Guide
 - Interactive Elements Periodic Trends Study Guide
- 8. Staying Engaged with Periodic Trends Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Study Guide
- 9. Balancing eBooks and Physical Books Periodic Trends Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Study Guide
 - Setting Reading Goals Periodic Trends Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Study Guide
 - Fact-Checking eBook Content of Periodic Trends Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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