

Periodic Trends Worksheet

Name: Grade: ID:

Directions: Use your notes to answer the following questions.

1 Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Ans: O, C, Al, K

2 Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Ans: Ne, Al, S, O

3 Why does fluorine have a higher ionization energy than iodine?

Ans: Fluorine has a smaller atomic radius, so the proton can exert a greater pull on electron (Valence shell is close)

4 Why do elements in the same family generally have similar properties?

Ans: They have the same valence shell electron arrangement.

5 Indicate whether the following properties increase or decrease from left to right across the periodic table.

a. atomic radius (excluding noble gases) **Decrease**

b. first ionization energy **Increase**

c. electronegativity **Increase**

6 What trend in atomic radius occurs across the periodic table? What causes this trend?

Ans: Atomic radius increases down a group because energy levels (shells) are added.

7 What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ans: Ionization energy increases across a period because as elements become less metallic, it requires more energy to remove an electron.

8 Circle the atom in each pair that has the largest radius.

a. **Al** or B

d. **O** or F

b. **Na** or Al

e. **Br** or Cl

c. **S** or O

f. Mg or **Ca**

Periodic Table Trends Answer Key

**Mark Weller, Mark T. Weller, Tina
Overton, Jonathan Rourke, Fraser
Armstrong**



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Chemistry in a Month John Wilkinson, 2003

Shriver and Atkins' Inorganic Chemistry Peter Atkins, 2010 Inorganic Chemistry fifth edition represents an integral part of a student s chemistry education Basic chemical principles are set out clearly in Foundations and are fully developed throughout the text culminating in the cutting edge research topics of the Frontiers which illustrate

the dynamic nature of inorganic chemistry

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Inorganic Chemistry Mark Weller, Mark T. Weller, Tina Overton, Jonathan Rourke, Fraser Armstrong, 2014 Leading the reader from the fundamental principles of inorganic chemistry right through to cutting edge research at the forefront of the subject Inorganic Chemistry Sixth Edition is the ideal course companion for the duration of a student's degree The authors have drawn upon their extensive teaching and research experience in updating this established text the sixth edition retains the much praised clarity of style and layout from previous editions while offering an enhanced Frontiers section Exciting new applications of inorganic chemistry have been added to this section in particular relating to materials chemistry and medicine This edition also sees a greater use of learning features to provide students with all the support they need for their studies Providing comprehensive coverage of inorganic chemistry while placing it in context this text will enable the reader to fully master this important subject Online Resource Centre For registered adopters of the text Figures marginal structures and tables of data ready to download Test bank For students Answers to self tests and exercises from the book Videos of chemical reactions Tables for group theory Web links Interactive structures and other resources on www.chemtube3d.com

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