

Name: **Brittany French**

Date: **October 4, 2012**

Student Exploration: Electron Configuration

Activity B: Atomic radii

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Na** (sodium).
- Select the ELECTRON CONFIGURATION tab.



Question: How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

It will decrease

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s² 2s² 2p⁶ 3s¹**

Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s² 2s² 2p⁶ 3s²**

Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s ² 2s ² 2p ⁶ 3s ² 3p ¹	118
Silicon	14	1s ² 2s ² 2p ⁶ 3s ² 3p ²	111
Phosphorus	15	1s ² 2s ² 2p ⁶ 3s ² 3p ³	98
Sulfur	16	1s ² 2s ² 2p ⁶ 3s ² 3p ⁴	88
Chlorine	17	1s ² 2s ² 2p ⁶ 3s ² 3p ⁵	79
Argon	18	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶	71

Periodic Table Exploration Answer Key

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Table of Contents Periodic Table Exploration Answer Key

1. Understanding the eBook Periodic Table Exploration Answer Key
 - The Rise of Digital Reading Periodic Table Exploration Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Exploration Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Exploration Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Exploration Answer Key
 - Personalized Recommendations
 - Periodic Table Exploration Answer Key User Reviews and Ratings
 - Periodic Table Exploration Answer Key and Bestseller Lists
5. Accessing Periodic Table Exploration Answer Key Free and Paid eBooks

- Periodic Table Exploration Answer Key Public Domain eBooks
 - Periodic Table Exploration Answer Key eBook Subscription Services
 - Periodic Table Exploration Answer Key Budget-Friendly Options
6. Navigating Periodic Table Exploration Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Table Exploration Answer Key Compatibility with Devices
 - Periodic Table Exploration Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Table Exploration Answer Key
 - Highlighting and Note-Taking Periodic Table Exploration Answer Key
 - Interactive Elements Periodic Table Exploration Answer Key
8. Staying Engaged with Periodic Table Exploration Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Exploration Answer Key
9. Balancing eBooks and Physical Books Periodic Table Exploration Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Exploration Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Exploration Answer Key
- Setting Reading Goals Periodic Table Exploration Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Exploration Answer Key
- Fact-Checking eBook Content of Periodic Table Exploration Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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