

# Periodic Table Basics

**Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.**

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

**Step 2: Determine the number of protons, neutrons, and electrons in each element.**

**Step 3: Create a Bohr diagram for each element.**

**Step 4: Draw the Lewis Structure for each element.**

The card for Boron (B) contains the following information:

- Top left: Atomic number 5
- Center: Element symbol B and name Boron
- Bottom left: Atomic mass 10.81
- Top right: Protons (P) = 5, Neutrons (N) = 6, Electrons (E) = 5
- Bottom right: Bohr Diagram (showing 5 electrons in two shells) and Lewis structure (B with three valence electrons represented by dots).

**Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.**

Green = Li & Na

Pink = O & S

Blue = Be & Mg

Purple = F & Cl

Orange = B & Al

Red = C & Si

Tan = N & P

Yellow = He, Ne, & Ar

**Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.**



**Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.**

# Periodic Table Basics Answer Key

**Matthew Johll**



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