

## CHAPTER 5 REVIEW

***The Periodic Law*****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- \_\_\_\_\_ In the modern periodic table, elements are ordered
  - \_\_\_\_\_ according to decreasing atomic mass.
  - \_\_\_\_\_ according to Mendeleev's original design.
  - \_\_\_\_\_ according to increasing atomic number.
  - \_\_\_\_\_ based on when they were discovered.
- \_\_\_\_\_ Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing
  - \_\_\_\_\_ density.
  - \_\_\_\_\_ reactivity.
  - \_\_\_\_\_ atomic number.
  - \_\_\_\_\_ atomic mass.
- \_\_\_\_\_ The modern periodic law states that
  - \_\_\_\_\_ no two electrons with the same spin can be found in the same place in an atom.
  - \_\_\_\_\_ the physical and chemical properties of an element are functions of its atomic number.
  - \_\_\_\_\_ electrons exhibit properties of both particles and waves.
  - \_\_\_\_\_ the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
- \_\_\_\_\_ The discovery of the noble gases changed Mendeleev's periodic table by adding a new
  - \_\_\_\_\_ period.
  - \_\_\_\_\_ series.
  - \_\_\_\_\_ group.
  - \_\_\_\_\_ level.
- \_\_\_\_\_ The most distinctive property of the noble gases is that they are
  - \_\_\_\_\_ metallic.
  - \_\_\_\_\_ radioactive.
  - \_\_\_\_\_ metalloid.
  - \_\_\_\_\_ largely unreactive.
- \_\_\_\_\_ Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
  - \_\_\_\_\_ 4.
  - \_\_\_\_\_ 10.
  - \_\_\_\_\_ 11.
  - \_\_\_\_\_ 18.
- \_\_\_\_\_ An isotope of fluorine has a mass number of 19 and an atomic number of 9.
  - \_\_\_\_\_ a. How many protons are in this atom?
  - \_\_\_\_\_ b. How many neutrons are in this atom?
  - \_\_\_\_\_ c. What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

# Periodic Law Answer Sheet

**Rachel Sandford**



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