

## Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

**Remember:** Limiting reagent is the reactant that will completely run out first in a chemical reaction.

**Excess:** All reagents except the limiting reagent will be in excess. Excess reagent is the reagent that is not completely consumed.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction. Excess reagent is the reagent that is not completely consumed.



The molar masses of the reactants and products are:

$\text{H}_2$	2.016 g/mol
$\text{O}_2$	31.998 g/mol
$\text{H}_2\text{O}$	18.015 g/mol
$\text{H}_2\text{O}$	18.015 g/mol

1. 10.00 g of  $\text{H}_2$  and 10.00 g of  $\text{O}_2$  are reacted. How much  $\text{H}_2\text{O}$  is produced? (10.00 g of  $\text{H}_2$  is the limiting reagent, 10.00 g of  $\text{O}_2$  is the excess reagent.)



2. 10.00 g of  $\text{H}_2$  and 10.00 g of  $\text{O}_2$  are reacted. How much  $\text{H}_2\text{O}$  is produced? (10.00 g of  $\text{H}_2$  is the limiting reagent, 10.00 g of  $\text{O}_2$  is the excess reagent.)

3. 10.00 g of  $\text{H}_2$  and 10.00 g of  $\text{O}_2$  are reacted. How much  $\text{H}_2\text{O}$  is produced? (10.00 g of  $\text{H}_2$  is the limiting reagent, 10.00 g of  $\text{O}_2$  is the excess reagent.)



4. 10.00 g of  $\text{H}_2$  and 10.00 g of  $\text{O}_2$  are reacted. How much  $\text{H}_2\text{O}$  is produced? (10.00 g of  $\text{H}_2$  is the limiting reagent, 10.00 g of  $\text{O}_2$  is the excess reagent.)



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**S Marginson**



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