

Relative Strengths of **Acids** and **Bases** in Aqueous Solutions at 25°C

Name of Acid	Acid		Base	K _a
Hydriodic	HI	→	H ⁺ + I ⁻	>> 1
Hydrobromic	HBr	→	H ⁺ + Br ⁻	>> 1
Hydrochloric	HCl	→	H ⁺ + Cl ⁻	>> 1
Sulfuric	H ₂ SO ₄	→	H ⁺ + HSO ₄ ⁻	> 1
Nitric	HNO ₃	→	H ⁺ + NO ₃ ⁻	> 1
Hydronium Ion	H ₃ O ⁺	⇌	H ⁺ + H ₂ O	5.50 X 10 ⁻¹
Oxalic	H ₂ C ₂ O ₄	⇌	H ⁺ + HC ₂ O ₄ ⁻	5.90 X 10 ⁻²
Sulfurous	H ₂ SO ₃	⇌	H ⁺ + HSO ₃ ⁻	1.41 X 10 ⁻²
Hydrogen Sulfate Ion	HSO ₄ ⁻	⇌	H ⁺ + SO ₄ ²⁻	1.20 X 10 ⁻²
Phosphoric	H ₃ PO ₄	⇌	H ⁺ + H ₂ PO ₄ ⁻	7.52 X 10 ⁻³
Nitrous	HNO ₂	⇌	H ⁺ + NO ₂ ⁻	5.62 X 10 ⁻⁴
Hydrofluoric	HF	⇌	H ⁺ + F ⁻	3.53 X 10 ⁻⁴
Formic	HCO ₂ H	⇌	H ⁺ + HCO ₂ ⁻	1.78 X 10 ⁻⁴
Benzoic	C ₆ H ₅ CO ₂ H	⇌	H ⁺ + C ₆ H ₅ CO ₂ ⁻	6.46 X 10 ⁻⁵
Acetic	CH ₃ CO ₂ H	⇌	H ⁺ + CH ₃ CO ₂ ⁻	1.76 X 10 ⁻⁵
Carbonic	H ₂ CO ₃	⇌	H ⁺ + HCO ₃ ⁻	4.30 X 10 ⁻⁷
Hydrogen Sulfite Ion	HSO ₃ ⁻	⇌	H ⁺ + SO ₃ ²⁻	1.02 X 10 ⁻⁷
Dihydrogen Phosphate Ion	H ₂ PO ₄ ⁻	⇌	H ⁺ + HPO ₄ ²⁻	6.23 X 10 ⁻⁸
Boric	H ₃ BO ₃	⇌	H ⁺ + H ₂ BO ₃ ⁻	5.79 X 10 ⁻¹⁰
Ammonium Ion	NH ₄ ⁺	⇌	H ⁺ + NH ₃	5.64 X 10 ⁻¹⁰
Hydrogen Carbonate Ion	HCO ₃ ⁻	⇌	H ⁺ + CO ₃ ²⁻	5.61 X 10 ⁻¹¹
Monohydrogen Phosphate Ion	HPO ₄ ²⁻	⇌	H ⁺ + PO ₄ ³⁻	2.20 X 10 ⁻¹³
Water	H ₂ O	⇌	H ⁺ + OH ⁻	1.00 X 10 ⁻¹⁴

Strong Acids

Weak Acids

Weakest Acids

Weakest Bases

Weak Bases

Stronger Bases

All reactions assume water as a reactant.



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$$K_a = [H^+][A^-] / [HA]$$

$$pK_a = -\log K_a$$

$$pH = -\log [H^+]$$

Section 2 Study Guid Strength Of Acid And Bases

Brendan G. Carr



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