ARCHO	Conscient			
2.	Number of atoms in 56 (1) Twice that in 70 g N (3) Both (1) and (2)		56 gmol-1) is : (2) Half that in 20 g H (4) None of these	
3.	Which has maximum n (1) 24 g of C (12)	(2) 56 g of Fe (56)	(3) 27 g of Al (27)	(4) 108 a Aa (108)
4.	The total number of pro (1) 1.084 = 10 ²⁵	otons, electrons and ne (2) 6.022 × 10 ²⁰	utrons in 12 g of #C is:	(4) 18
5.0	1.0 g of hydrogen contains 6×10 ²⁰ atoms. The atomic weight of helium is 4, it follows that the number of atoms in 1 g of He is :			
	(1) $\frac{1}{4} \times 6 \times 10^{20}$	$(2) 4 \times 6 \times 10^{23}$	(3) 6 × 10 ²³	(4) 12 × 10 ²³
	A sample of aluminium atoms? (At. wt. Al = 2)	m has a mass of 54.0 7, Mg = 24) (2) 24 o	g. What is the mass of t	he same number of magnesiu (4) 96 g.
7.	The number of atoms in 558.5 g of Fe (at wt. = 55.85) is : (1) Twice that in 60 g carbon (2) 6.022 × 10 ²² (3) Half in 8 g He (4) 558.5 × 6.023 × 10 ²³			
8.	The total number of e	lectrons present in 8.0 (2) 3.01 × 10 ²⁴	g of methane is (3) 4.8 × 10 ²⁵	(4) 3.01 × 10 ²³ .
9.	How many moles of e	lectron weigh one kilog	gram :	221
	(1) 6.023 × 10 ²³			$(4) \ \frac{1}{9.108 \times 6.023} \times 10^{6}$
0.	Under the same conditions, two gases have the same number of molecules. They must (1) be noble gases (3) have a volume of 22.4 dm³ each (4) have an equal number of atoms			
1.	(1) 6.02 × 10 ²³ atoms	of oxygen at STP contains : (2) 3.01×10^{23} atoms (4) 0.7525×10^{23} atoms (4) 0.7525×10^{23} atoms of the vapours of substance at NTP weight W g. Then molecular weight of substance is : $W \times 1$		
2.	If V ml of the vapours	The second secon	weight W g. Then molecul	lar weight of substance is:
	(1) (W/V) × 22400	(2) $\frac{V}{W} \times 22.4$	(3) (W-V) × 22400	(4) V×22400
3.	Control of the second		ram-molecular weight is : (3) 22.4 g	(4) 224 g
24.	(1) x = 3	(2) 4 - 2	STP. The value of x is (3) x = 4	(4) none of these
Sec	tion (C) : Density ar	nd Vapour density	/	
1.	Vapour density of a	(2) 2	(3) 4	(4) 0.089
2.	(1) 7	(2) 4	t's VD is 70 the value of (3) 5	x must be : (4) 6
3.	$(11.6.02 \times 10^{23})$	in 1.8 mL of H ₂ O(I) is (2) 3.011 × 10 ²³	(3) 0.0000	
Sec	tion (D) : Percentag	e composition ar	nd Molecular formu	ila
	The percentage value	ie of nitrogen in urea i	s about (3) 18	(4) 28

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Camilla Rothe

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