

14.3

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Calculate concentrations based on the pH of a given acid or base solution and vice versa.
2. Calculate pH values and concentrations of acids and bases and vice versa.
3. Calculate the pH of a buffer (see Section 14.4).

Vocabulary

- | | | |
|-------------------|---|-----------------------|
| 1. autoionization | 2. hydronium ion (H_3O^+) | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|---|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |

Part A Completion

Read the paragraph and use the information to determine if the sentence is true or false. Write true or false in the space provided. If the sentence is false, rewrite it so that it is a true statement.

1. When a substance ionizes, it releases hydrogen ions into the solution. (true)
2. The concentration of H^+ in the autoionization of water is 1.0×10^{-7} moles per liter at 25°C. (true)
3. The concentration of H^+ in a 0.1 M solution of HCl is 1.0×10^{-7} moles per liter. (false) 1.0×10^{-1} moles per liter
4. The pH of a 0.1 M solution of HCl is 1. (true)
5. The pH of a 0.1 M solution of HCl is 1. (true)
6. The pH of a 0.1 M solution of HCl is 1. (true)
7. The pH of a 0.1 M solution of HCl is 1. (true)
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16. The pH of a 0.1 M solution of HCl is 1. (true)
17. The pH of a 0.1 M solution of HCl is 1. (true)
18. The pH of a 0.1 M solution of HCl is 1. (true)
19. The pH of a 0.1 M solution of HCl is 1. (true)
20. The pH of a 0.1 M solution of HCl is 1. (true)

1. $\text{pH} = -\log [\text{H}^+]$
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Section 19 2 Hydrogen Ions And Acidity Answer Key

Lei Huang



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