Limiting Reactant and Percentage Yield

1 Calcium hydroxide reacts with hydrochloric acid according to the following reaction:

$$Ca(OH)_2 + 2\ HCl \longrightarrow CaCl_2 + 2\ H_2O$$

(a) Suppose 6.3 mol of HCl reacts with 2.8 mol $Ca(OH)_2$. Which substance is the limiting reactant?

(b) How many moles of the excess reactant remain?

2 Aluminum oxidizes according to the following unbalanced reaction: (First, balance it)

$$\_\_\_ Al + \_\_\_ O_2 \longrightarrow \_\_\_ Al_2O_3$$

(a) 0.048 mol of powdered Al is placed in a container containing 0.030 mol of O_2. What is the limiting reactant?

(b) How many moles of excess reactant remain?

3 In the reaction between CO and Fe_3O_4, the theoretical yield in an experiment is calculated to be 47.2 g Fe. When a careless chemistry student experiments, the actual yield is 42.9 g Fe. Calculate the percentage yield.

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