

Chapter 10 – CHEMICAL QUANTITIES

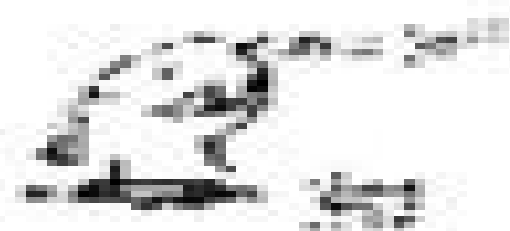
The MOLE

Avogadro's Hypothesis

Equal volumes of gases at constant T and p contain the same # molecules.

The number of H_2 atoms in 12.01 grams of carbon is called **Avogadro's Number** – 6.02×10^{23}

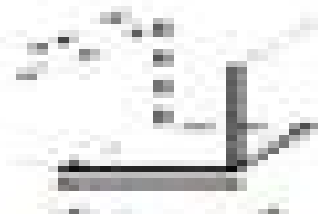
This quantity is called a **MOLE**.
(Just one concept = 12)



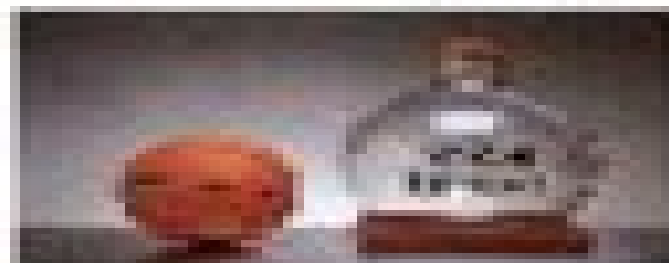
1 mole = 6.02×10^{23} particles (approx. number of molecules)

STP: Standard Temperature and Pressure
0°C 1 atm

At STP, **1 mole = 22.4 L** for any gas.



At STP, 1 mole is **22.4 L**,
approximately 22.4 L at any temp.



Section 10 Chemical Quantities Packet Answers

Michael Rajnik



Section 10 Chemical Quantities Packet Answers:

Section 10 Chemical Quantities Packet Answers Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the ability of words has become more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Section 10 Chemical Quantities Packet Answers**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall affect readers.

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