

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance
 - Volume of an object? ruler or water displacement method
 - Length of an object? Ruler
 - Density of an object? graduated cylinder
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8.33/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

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D Kirk



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Table of Contents Second Semester Final Study Guide

1. Understanding the eBook Second Semester Final Study Guide
 - The Rise of Digital Reading Second Semester Final Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Second Semester Final Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Second Semester Final Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Second Semester Final Study Guide
 - Personalized Recommendations
 - Second Semester Final Study Guide User Reviews and Ratings
 - Second Semester Final Study Guide and Bestseller Lists
5. Accessing Second Semester Final Study Guide Free and Paid eBooks
 - Second Semester Final Study Guide Public Domain eBooks
 - Second Semester Final Study Guide eBook Subscription Services
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6. Navigating Second Semester Final Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Second Semester Final Study Guide Compatibility with Devices
 - Second Semester Final Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Second Semester Final Study Guide
 - Highlighting and Note-Taking Second Semester Final Study Guide
 - Interactive Elements Second Semester Final Study Guide

8. Staying Engaged with Second Semester Final Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Second Semester Final Study Guide
9. Balancing eBooks and Physical Books Second Semester Final Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Second Semester Final Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Second Semester Final Study Guide
 - Setting Reading Goals Second Semester Final Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Second Semester Final Study Guide
 - Fact-Checking eBook Content of Second Semester Final Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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