

Sample Test Paper - II

Course Name : Civil Engineering Group

Course Code : CE/CS/CR/ CV

Semester : Second

Subject Title : Applied Science (Physics)

Marks : 25

17207

Time:1 Hour

Instructions:

1. All questions are compulsory.
2. Illustrate your answers with neat sketches wherever necessary.
3. Figures to the right indicate full marks.
4. Assume suitable data if necessary.
5. Preferably, write the answers in sequential order.

Q.1) Answer any FOUR

08 Marks

- a. Define echo and reverberation.
- b. State Inverse square law.
- c. State Sabine's formula with symbol meaning.
- d. State any two properties of photon.
- e. State any two requirements of good acoustics.
- f. An X ray tube works on 40 kv. What will be the wavelength of X rays emitted in it.

Q.2) Answer any THREE

09 Marks

- a. Explain Planck's hypothesis.
- b. State any six applications of X rays.
- c. A concrete hall of volume 2000 m^3 has total absorption of 200. Find the reverberation time.
- d. Two sources of equal illumination power are placed at a distance of 1.8 m from one another. Where should a screen be placed between two sources so that the intensity of illumination on one of the surface be 4 times on the other.

Q.3) Answer any TWO

08 Marks

- a. State four factors affecting indoor lighting system and explain any two.
- b. Explain production of X rays using Coolidge's tube.
- c. The photo electric work function of certain metal is 3×10^{-19} joules. Calculate it's threshold frequency if Planck's constant is $6.625 \times 10^{-34} \text{ Js}$

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G Psacharopoulos



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Table of Contents Sample Question Paper Diploma Civil Engineering Msbte

1. Understanding the eBook Sample Question Paper Diploma Civil Engineering Msbte
 - The Rise of Digital Reading Sample Question Paper Diploma Civil Engineering Msbte
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample Question Paper Diploma Civil Engineering Msbte
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Question Paper Diploma Civil Engineering Msbte
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Question Paper Diploma Civil Engineering Msbte
 - Personalized Recommendations
 - Sample Question Paper Diploma Civil Engineering Msbte User Reviews and Ratings
 - Sample Question Paper Diploma Civil Engineering Msbte and Bestseller Lists
5. Accessing Sample Question Paper Diploma Civil Engineering Msbte Free and Paid eBooks
 - Sample Question Paper Diploma Civil Engineering Msbte Public Domain eBooks
 - Sample Question Paper Diploma Civil Engineering Msbte eBook Subscription Services
 - Sample Question Paper Diploma Civil Engineering Msbte Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Sample Question Paper Diploma Civil Engineering Msbte Compatibility with Devices
 - Sample Question Paper Diploma Civil Engineering Msbte Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Question Paper Diploma Civil Engineering Msbte
 - Highlighting and Note-Taking Sample Question Paper Diploma Civil Engineering Msbte
 - Interactive Elements Sample Question Paper Diploma Civil Engineering Msbte
8. Staying Engaged with Sample Question Paper Diploma Civil Engineering Msbte
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample Question Paper Diploma Civil Engineering Msbte
9. Balancing eBooks and Physical Books Sample Question Paper Diploma Civil Engineering Msbte
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Question Paper Diploma Civil Engineering Msbte
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Sample Question Paper Diploma Civil Engineering Msbte
 - Setting Reading Goals Sample Question Paper Diploma Civil Engineering Msbte
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sample Question Paper Diploma Civil Engineering Msbte
 - Fact-Checking eBook Content of Sample Question Paper Diploma Civil Engineering Msbte
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1250 \text{ J}$
 $550 \text{ J} = \Delta h \cdot q_{\text{water}} = 550 \text{ J}$

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 for water $4.184 \text{ J/g} \cdot \text{C}$ 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by 8.1 C when 207 J of heat is added 2

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 25.0 C and heat it to 145.0 C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the
 process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 C

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 of 1.00 M NaOH at 25.0 C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0 C calculate the
 enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/ml and that

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web 6 a 218 g sample of steam at 121 $^{\circ}\text{C}$ is cooled to ice at 14 $^{\circ}\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^{\circ}\text{C}$ to steam at 138 $^{\circ}\text{C}$ a sample of water absorbs $1.41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 $\times 10^4 \text{ J}$ 2 6 23 $\times 10^4 \text{ J}$ 3 1 71 $\times 10^4 \text{ J}$ 4 1 11 $\times 10^5 \text{ J}$ 5 2 11 $\times 10^5 \text{ J}$ 6

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