

Name: Key

Class: _____

Date: _____

Stoichiometry Problem Set A

Directions: Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

Part 1: Straight Stoichiometry Problems

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$, how many grams of magnesium ^B nitrate are produced from 8.00 mol of nitric acid (HNO_3)? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$, how many moles of potassium ^B chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{mole KClO}_3$$

- 3.) For the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of water are produced from 5.22 liters of oxygen gas? ^B A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of hydrogen ^B are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$, how many moles of carbonic ^B acid can be produced from 97.2 g of CO_2 ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{moles H}_2\text{CO}_3$$

Did you achieve our target goal?

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1
Feeling VERY confused

2
Not sure that I get this

3
I think I am catching on

4
Doing great! Bring it on!

Sample Problem Set Stoichiometry Chemistry

RM Cervero

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

Sample Problem Set Stoichiometry Chemistry:

Adopting the Melody of Appearance: An Mental Symphony within **Sample Problem Set Stoichiometry Chemistry**

In some sort of used by monitors and the ceaseless chatter of immediate interaction, the melodic splendor and mental symphony created by the prepared term frequently diminish in to the background, eclipsed by the constant noise and distractions that permeate our lives. Nevertheless, set within the pages of **Sample Problem Set Stoichiometry Chemistry** an enchanting literary treasure overflowing with fresh feelings, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, this charming masterpiece conducts visitors on a mental journey, well unraveling the hidden tunes and profound affect resonating within each cautiously crafted phrase. Within the depths with this touching assessment, we shall investigate the book is central harmonies, analyze their enthralling writing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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Table of Contents Sample Problem Set Stoichiometry Chemistry

1. Understanding the eBook Sample Problem Set Stoichiometry Chemistry
 - The Rise of Digital Reading Sample Problem Set Stoichiometry Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample Problem Set Stoichiometry Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Problem Set Stoichiometry Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Problem Set Stoichiometry Chemistry
 - Personalized Recommendations

- Sample Problem Set Stoichiometry Chemistry User Reviews and Ratings
- Sample Problem Set Stoichiometry Chemistry and Bestseller Lists
- 5. Accessing Sample Problem Set Stoichiometry Chemistry Free and Paid eBooks
 - Sample Problem Set Stoichiometry Chemistry Public Domain eBooks
 - Sample Problem Set Stoichiometry Chemistry eBook Subscription Services
 - Sample Problem Set Stoichiometry Chemistry Budget-Friendly Options
- 6. Navigating Sample Problem Set Stoichiometry Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Sample Problem Set Stoichiometry Chemistry Compatibility with Devices
 - Sample Problem Set Stoichiometry Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Problem Set Stoichiometry Chemistry
 - Highlighting and Note-Taking Sample Problem Set Stoichiometry Chemistry
 - Interactive Elements Sample Problem Set Stoichiometry Chemistry
- 8. Staying Engaged with Sample Problem Set Stoichiometry Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample Problem Set Stoichiometry Chemistry
- 9. Balancing eBooks and Physical Books Sample Problem Set Stoichiometry Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Problem Set Stoichiometry Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Sample Problem Set Stoichiometry Chemistry
 - Setting Reading Goals Sample Problem Set Stoichiometry Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sample Problem Set Stoichiometry Chemistry
 - Fact-Checking eBook Content of Sample Problem Set Stoichiometry Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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