# Percent Composition and Molecular Formula Worksheet Key

- What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? C<sub>3</sub>H<sub>3</sub>O mass = 55 g/mole
- If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula?
   C<sub>6</sub>H<sub>6</sub>O<sub>2</sub>
- What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? Li<sub>2</sub>CO<sub>3</sub>
- If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula?
   Li<sub>2</sub>CO<sub>3</sub>

#### Write the molecular formulas of the following compounds:

- A compound with an empirical formula of C<sub>2</sub>OH<sub>4</sub> and a molar mass of 88 grams per mole. C<sub>4</sub>O<sub>2</sub>H<sub>8</sub>
- 6. A compound with an empirical formula of C4H4O and a molar mass of 136 grams per mole. C8H8O2
- A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole. C<sub>2</sub>F<sub>2</sub>Br<sub>2</sub>O<sub>2</sub>
- 8. A compound with an empirical formula of C2H8N and a molar mass of 46 grams per mole. C2H8N

#### Answer the following questions:

- The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. CH<sub>2</sub>O
- 10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>
- 11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO<sub>2</sub>, H<sub>2</sub>O, and N<sub>2</sub> as products. If the combustion of 9.71 g of aniline yields 6.63 g H<sub>2</sub>O and 1.46 g N<sub>2</sub>, what is its empirical formula? C<sub>6</sub>H<sub>7</sub>N
- 12. The molar mass of aniline is 93 g/mol. What is its molecular formula? CoH1N
- Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C<sub>2</sub>H<sub>5</sub>NO. %C 40.668 %H 8.533 %N 23.713 %O 27.086
- 14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? PCIs
- When 2.5000 g of an oxide of mercury, (Hg<sub>x</sub>O<sub>y</sub>) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. Hg<sub>2</sub>O
- 16. The compound benzamide has the following percent composition. What is the empirical formula?
  C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % C<sub>7</sub>H<sub>7</sub>NO
- 17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?

C<sub>3</sub>H<sub>2</sub>NO<sub>3</sub> empirical formula C<sub>4</sub>H<sub>2</sub>NO<sub>3</sub> molecular formula

# **Percent Composition And Molecular Formula Answers**

**ML Morrison** 

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