



Lab Notebook

Classifying Chemical Reactions

Directions: Write the balanced chemical equation for each reaction. Classify each reaction as one of the following: synthesis, decomposition, single displacement, double displacement, combustion, or other.

Column I

1. A. In precipitation reactions, solid is a gas forms collected from two aqueous solutions (precipitate)
2. B. Two or more substances combine to form a single substance
3. C. A compound breaks down into two or more simpler substances
4. D. One element replaces another in a compound
5. E. A type of redox reaction that produces heat and light

Directions: Write the balanced chemical equation for each reaction. Classify each reaction as one of the following: synthesis, decomposition, single displacement, double displacement, combustion, or other.

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

Column II

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

Column III

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

Column IV

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

Column V

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

Column VI

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

Column VII

1. A. $2H_2 + O_2 \rightarrow 2H_2O$
2. B. $2H_2O \rightarrow 2H_2 + O_2$

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