

## Behavior of Gases

force  
partides  
Boyle's  
constantly

increase  
Charles's  
absolute  
liquids

kinetic  
size  
decrease  
pressure

pressure  
boiling  
larger  
decrease

volume  
kilopascals  
temperature  
increased

Gases in Earth's atmosphere exert 1. \_\_\_\_\_ on everything. According to the 2. \_\_\_\_\_ theory, the particles of a gas are 3. \_\_\_\_\_ moving. Every time gas particles hit something and bounce off, they exert a tiny force. Pressure is this amount of 4. \_\_\_\_\_ exerted per unit of area. Air pressure at sea level is 101.3 5. \_\_\_\_\_.

The amount of force exerted by a gas depends on the 6. \_\_\_\_\_ of its container. 7. \_\_\_\_\_ law states that if a sample of gas is kept at constant 8. \_\_\_\_\_, decreasing the volume will 9. \_\_\_\_\_ the pressure the gas exerts. If you increase the volume, the pressure will 10. \_\_\_\_\_. According to the kinetic theory, if you do not change the amount of gas or its temperature but 11. \_\_\_\_\_ the size of the container, the particles will strike the walls more often and the pressure will rise. When the size of the container is 12. \_\_\_\_\_, the pressure is smaller because the 13. \_\_\_\_\_ hit the walls less often.

According to 14. \_\_\_\_\_ law, if a sample of gas is kept at constant 15. \_\_\_\_\_, the volume increases if the temperature is 16. \_\_\_\_\_. Charles's measurements suggested that the 17. \_\_\_\_\_ of a gas would become zero at a temperature of  $-273^{\circ}\text{C}$ . The temperature  $-273^{\circ}\text{C}$  is called 18. \_\_\_\_\_ zero. All gases become 19. \_\_\_\_\_ when cooled to their 20. \_\_\_\_\_ points.

# Section 13 1 Gases Chemistry Worksheet Awnsers

**Lingjun Ying**



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