

11

CHEMICAL REACTIONS

Practice Problems

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

1. Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
2. Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
3. Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
4. Write a balanced equation for the production of HCl gas from its elements.
5. Write a sentence that completely describes the chemical reaction represented by this balanced equation.



6. Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.



7. Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
8. Write a balanced equation from this word equation.
aqueous silver nitrate + copper metal $\rightarrow$ silver metal + aqueous copper nitrate
9. Write a balanced equation for the following word equation.



SECTION 11.2 TYPES OF CHEMICAL REACTIONS

1. Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
2. Write the balanced equation for the reaction that occurs between aluminum and fluorine.
3. Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
4. Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
5. Write the balanced equation for the combustion of propane (C_3H_8) to produce carbon dioxide and water vapor.
6. Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

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Ralph Thomas Myers



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Table of Contents Section 11 1 Describing Chemical Reaction Answer Key

1. Understanding the eBook Section 11 1 Describing Chemical Reaction Answer Key
 - The Rise of Digital Reading Section 11 1 Describing Chemical Reaction Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 11 1 Describing Chemical Reaction Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 11 1 Describing Chemical Reaction Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 11 1 Describing Chemical Reaction Answer Key
 - Personalized Recommendations
 - Section 11 1 Describing Chemical Reaction Answer Key User Reviews and Ratings
 - Section 11 1 Describing Chemical Reaction Answer Key and Bestseller Lists
5. Accessing Section 11 1 Describing Chemical Reaction Answer Key Free and Paid eBooks
 - Section 11 1 Describing Chemical Reaction Answer Key Public Domain eBooks
 - Section 11 1 Describing Chemical Reaction Answer Key eBook Subscription Services
 - Section 11 1 Describing Chemical Reaction Answer Key Budget-Friendly Options
6. Navigating Section 11 1 Describing Chemical Reaction Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Section 11 1 Describing Chemical Reaction Answer Key Compatibility with Devices
 - Section 11 1 Describing Chemical Reaction Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Section 11 1 Describing Chemical Reaction Answer Key
 - Highlighting and Note-Taking Section 11 1 Describing Chemical Reaction Answer Key
 - Interactive Elements Section 11 1 Describing Chemical Reaction Answer Key
8. Staying Engaged with Section 11 1 Describing Chemical Reaction Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 11 1 Describing Chemical Reaction Answer Key
9. Balancing eBooks and Physical Books Section 11 1 Describing Chemical Reaction Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 11 1 Describing Chemical Reaction Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 11 1 Describing Chemical Reaction Answer Key
- Setting Reading Goals Section 11 1 Describing Chemical Reaction Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 11 1 Describing Chemical Reaction Answer Key
- Fact-Checking eBook Content of Section 11 1 Describing Chemical Reaction Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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