

1. Show your answer to the following equation as either the reactant or product:



The particle represented by  $X$  is:

- A)  ${}^1_0\text{H}$  B)  ${}^4_2\text{He}$  C)  ${}^4_2\text{He}$  D)  ${}^1_0\text{H}$

2. Which isotope will spontaneously decay and emit particles with a charge of  $+2$ ?

- A)  ${}^{238}_{92}\text{U}$  B)  ${}^{137}_{54}\text{Cs}$  C)  ${}^{235}_{92}\text{U}$  D)  ${}^{239}_{94}\text{Pu}$

3. Which of these particles has the greatest mass?

- A) alpha B) beta  
C) neutron D) positron

4. Show your answer to the following equation as either the reactant or product:



What particle is represented by  $X$ ?

- A) alpha B) beta  
C) neutron D) positron

5. Alpha particles and beta particles differ in:

- A) mass, only  
B) charge, only  
C) both mass and charge  
D) neither mass nor charge

6. Which statement best describes gamma radiation?

- A) It has a mass of 1 and a charge of 1.  
B) It has a mass of 0 and a charge of  $-1$ .  
C) It has a mass of 0 and a charge of 0.  
D) It has a mass of 4 and a charge of  $+2$ .

7. Which type of radiation is most similar to high-energy x-rays?

- A) alpha B) beta  
C) neutron D) gamma

8. What is the total number of years that must pass before only 25.00 grams of an original 100.0-gram sample of  ${}^{60}_{27}\text{Co}$  remains unchanged?

- A) 1860 y B) 5160 y  
C) 11 460 y D) 17 190 y

9. Which isotope is most commonly used in the radioactive dating of the remains of organic materials?

- A)  ${}^3\text{H}$  B)  ${}^{14}\text{C}$  C)  ${}^{15}\text{N}$  D)  ${}^{18}\text{O}$

10. Based on Reference Table N, what fraction of a radioactive  ${}^{90}\text{Sr}$  sample would remain unchanged after 56.2 years?

- A)  $\frac{1}{2}$  B)  $\frac{1}{4}$  C)  $\frac{1}{8}$  D)  $\frac{1}{16}$

11. What is the half-life and decay mode of  $\text{Ra-226}$ ?

- A) 1.91 days and alpha decay  
B) 1.91 days and beta decay  
C) 1.62 days and alpha decay  
D) 1.62 days and beta decay

12. After 32 days, 3 milligrams of an 80-milligram sample of a radioactive isotope remains unchanged. What is the half-life of this element?

- A) 8 days B) 2 days  
C) 16 days D) 4 days

13. Which reaction converts an atom of one element to an atom of another element?

- A) combination B) polymerization  
C) saponification D) transmutation

14. Radioactive cobalt-60 is used in radiation therapy treatment. Cobalt-60 undergoes beta decay. This type of nuclear reaction is called:

- A) natural transmutation  
B) artificial transmutation  
C) nuclear fusion  
D) nuclear fission

15. In which reaction is mass converted to energy by the process of fusion?

- A)  ${}^1_0\text{H} + {}^1_0\text{H} \rightarrow {}^2_1\text{H} + {}^1_1\text{H}$   
B)  ${}^{235}_{92}\text{U} + {}^1_0\text{n} \rightarrow {}^{141}_{54}\text{Ba} + {}^{89}_{36}\text{Kr} + 3 {}^1_0\text{n}$   
C)  ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + {}^4_2\text{He}$   
D)  ${}^1_0\text{H} + {}^1_1\text{H} \rightarrow {}^2_1\text{H}$

16. Which statement best describes what happens in a fission reaction?

- A) Heavy nuclei split into lighter nuclei.  
B) Light nuclei form into heavier nuclei.  
C) Energy is released and less stable elements are formed.  
D) Energy is absorbed and more stable elements are formed.

# Section 24 Nuclear Chemistry Study Guide Key

**Robert J. Ouellette, J. David Rawn**



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