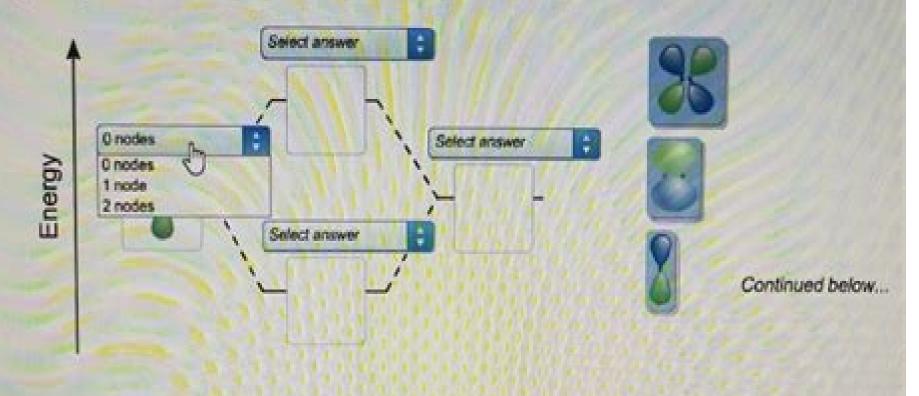
a) Below, construct an orbital interaction diagram for molecular orbital formation by dragging the images that represent various orbital types (e.g., atomic, bonding, antibonding) into the relevant boxes.

b) Identify the number of nodes in each atomic and molecular orbital.



c) When two electrons occupy the bonding molecular orbital above, what type of bond results? Explain.

- a because only one bond forms.
- ir because the bond is not cylindrically symmetrical about the internuclear axis,
- σ because the two electrons are in a σ bonding molecular orbital.
- π because p orbitals can never form σ bonds.

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