

DESCRIBING CHEMICAL REACTIONS

Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation

Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical ____1____. The substances that undergo a chemical change are the ____2____. The new substances formed in a chemical reaction are the ____3____. In accordance with the law of conservation of ____4____, a chemical equation must be balanced. When balancing an equation, you place ____5____ in front of reactants and products so that the same number of atoms of each ____6____ are on each side of the equation. An equation must never be balanced by changing the ____7____ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is ____8____; for a solid, ____9____; and for a gas, ____10____. A substance dissolved in water is designated ____11____. If a ____12____ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- _____ 14. The symbol Δ placed over the arrow in an equation means that heat is supplied to the reaction.
- _____ 15. Atoms are destroyed in a chemical reaction.
- _____ 16. A skeleton equation is not a balanced equation.

Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I
- Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.
 - Acetylene gas (C_2H_2) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

Section 2 Describing Chemical Reactions Answer Sheet

ML Yell



Section 2 Describing Chemical Reactions Answer Sheet:

Eventually, you will utterly discover a other experience and triumph by spending more cash. nevertheless when? pull off you tolerate that you require to acquire those all needs following having significantly cash? Why dont you attempt to get something basic in the beginning? Thats something that will lead you to understand even more in this area the globe, experience, some places, in the manner of history, amusement, and a lot more?

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