



This is the expression for the equilibrium constant for the reaction shown above.

☐ A $K_{\text{eq}} = [\text{CO}_2][\text{H}_2]$

☐ B $K_{\text{eq}} = \frac{[\text{CO}_2]^2[\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]^2}$

☐ C $K_{\text{eq}} = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$

☐ D $K_{\text{eq}} = \frac{[\text{CO}][\text{H}_2\text{O}]}{[\text{CO}_2][\text{H}_2]}$

2. The hydrogen sulfide produced as a byproduct of petroleum refinement can be used to produce elemental sulfur:



What is the equilibrium constant expression for this reaction?

☐ A $K_{\text{eq}} = \frac{[\text{H}_2\text{O}]^2}{[\text{H}_2\text{S}]^2[\text{SO}_2]}$

☐ B $K_{\text{eq}} = \frac{[\text{S}]^3[\text{H}_2\text{O}]^2}{[\text{H}_2\text{S}]^2[\text{SO}_2]}$

☐ C $K_{\text{eq}} = \frac{[\text{H}_2\text{S}]^2[\text{SO}_2]}{[\text{H}_2\text{S}]^2}$

☐ D $K_{\text{eq}} = \frac{[\text{H}_2\text{O}]}{[\text{H}_2\text{S}][\text{SO}_2]}$

3. What can be determined if the ion product (Q_{sp}) is compared to the solubility product constant (K_{sp}) ?

☐ A the law of mass action

☐ B whether a substance will precipitate

☐ C equilibrium

☐ D the volume of the solution

4. The value of any equilibrium constant is correct only at

☐ A a specific volume.

☐ B a specific pressure.

☐ C a specific concentration.

☐ D a specific temperature.

5. A system reaches chemical equilibrium when

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Answer Sheet

RM Cervero



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