

Practice Problems

Do your best to solve the following problems.

SECTION 12.1 THE MOLE: A MEASUREMENT OF MATTER

1. Calculate the molar mass of each of the following compounds:

a. sodium chloride (table salt) or sodium hypochlorite

b. propylene glycol (antifreeze) or ethylene glycol

c. aluminum sulfate (antacid) or aluminum

2. Calculate the mass percent of each of the following compounds:

a. sodium chloride

b. propylene glycol

3. Determine the molar mass of each of the following compounds:

4. Determine the mass percent of each of the following compounds:

5. Determine the mass of 1 mol of each of the following:

SECTION 12.2 MOLE—MASS AND MOLE—VOLUME RELATIONSHIPS

6. Calculate the molar mass of each of the following compounds:

a. H_2O , H_2O_2 , H_2SO_4

b. NaOH , Na_2O

c. CO_2 , CO

d. HNO_3 , HNO_2

7. Calculate the mass in grams of each of the following:

a. 0.5 mol of sodium chloride (table salt)

b. 0.01 mol of sodium hypochlorite (bleach)

c. 100 g of sodium sulfate (antacid)

d. 0.001 mol of sodium hypochlorite (bleach)

e. 0.0001 mol of sodium hypochlorite (bleach)

8. Determine the mass of each of the following compounds: 0.001 mol of

9. Determine the mass of each of the following compounds: 0.001 mol of

10. Determine the mass of each of the following compounds:

a. 0.01 mol of sodium chloride (table salt) or 0.01 mol of sodium hypochlorite (bleach)

b. 0.001 mol of sodium hypochlorite (bleach) or 0.001 mol of sodium sulfate (antacid)

c. 0.0001 mol of sodium hypochlorite (bleach)

11. Calculate the mass of each of the following compounds: 0.001 mol of

12. Calculate the mass of each of the following compounds: 0.001 mol of

Section 10 Chemical Quantities Practice Problem Answers

M Carnoy



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