#### Chemistry -- Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

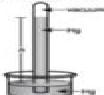
#### Multiple Choice

Identify the choice that best completes the statement or answers the question.

At 22°C, the air pressure in a car tire is 1293 torr.

Convert this pressure to atmospheres (atm).

- a. 12.76
- b. 5.262
- c. 1.701
- d. 1.293
- 2. This diagram shows a closed-end mercury (Hg)



barometer.

When atmospheric pressure is 1.30 atm, what is the height (h), in millimeters, of the Hg in the barometer?

- a. 132
- b. 585
- c. 760
- d. 988
- A chemist compresses a 1.5 L sample of carbon dioxide (CO<sub>2</sub>) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the CO<sub>2</sub> gas?
  - a. 0.60
  - b. 0.94
  - c. 1.7
  - d. 2.4
- A vessel connected to an open-end mercury (Hg)
  manometer contains nitrogen (N<sub>2</sub>) gas. The
  atmospheric pressure is 0.993 atm. The height
  difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the N2 gas?

- a. 355
- b. 633

- c. 755
- d. 877
- At 1.00 atm, a scaled weather balloon contains 20.0 L of helium (He) gas at 25.0°C. Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at 35.0°C?
  - a. 40.0
  - b. 28.0
  - c. 20.7
  - d. 14.3
- 6. At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at 95.9°C. Which statement explains the student's results at this high elevation?
  - The boiling point of water increases due to lower atmospheric pressure.
  - The boiling point of water increases due to higher atmospheric pressure.
  - The boiling point of water decreases due to lower atmospheric pressure.
  - The boiling point of water decreases due to higher atmospheric pressure.
- For a chemistry homework assignment, Beth must consider the effusion rates of methane (CH<sub>4</sub>) gas and nitrogen trifluoride (NF<sub>2</sub>) gas at 25°C. Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of CH<sub>4</sub> and NF<sub>3</sub> at 25°C?
  - a. CH<sub>4</sub> effuses 2 times faster than NF<sub>3</sub>.
  - b. CH<sub>4</sub> effuses 4 times faster than NF<sub>3</sub>.
  - c. NF3 effuses 2 times faster than CH4.
  - d. NF<sub>3</sub> effuses 4 times faster than CH<sub>4</sub>.
- At standard temperature and pessure, a chemist burns 3.50 L of pentane (C<sub>2</sub>H<sub>12</sub>) in excess oxygen (O<sub>2</sub>) gas. C<sub>2</sub>H<sub>12</sub>(I) + 8 O<sub>2</sub>(g) ⇒ 5 CO<sub>2</sub>(g) + 6 H<sub>2</sub>O(I) What is the maximum number of liters of carbon dioxide (CO<sub>2</sub>) that the reaction can produce?
  - a. 17.5
  - b. 8,50
  - c. 5.60
  - d. 0.700

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