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Arthur E. Morton



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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it s written like this
$$\frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100$$
 end array

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