

**Introduction:**

The laboratory method used in the experiment is titration. Titration is a method used in measuring the amount of an analytical reagent necessary to react quantitatively with the sample. Acid-base titrations are important for counting concentrations of acids and bases. The point at which the equivalence point has been reached is called the theoretical end point. The titration is detected by some physical change, such as colour change, to indicate the end point. The indicator solution used in this experiment is phenolphthalein which is a colourless in acidic solution. When the reaction reached the endpoint, the solution becomes pale pink. One or two drops of phenolphthalein are sufficient for the titration. Addition of too much indicator will necessitate an indicator blank. The indicator should give a clear visual change in the liquid being titrated once the reaction between the standard solution and the substance is practically complete. The indicators can change colour because their ions have colours that are different from the undissociates molecule (Tan, 2010). The acid and base titrations used the Arrhenius theory. This theory stated that acid are substance which produce hydrogen ions in solution and bases are substance which produce hydroxide ion in solution. In titration, standard solutions are now usually expressed in terms of molarity, M.

Molarity, M = moles of solute/volume of solutions in litres

**Objective :**

One of the purposes of this experiment is to demonstrate the basic laboratory technique of titration towards the students. This experiment enables a hand-on practice done by the students themselves. Thus, the students will know how to handle the materials, apparatus and the correct way to conduct the experiment as well as being careful to avoid any experimental error. Moreover, this experiment also teach the students on the correct steps to calculate percentage of substance, molarity of solution and other calculations involve throughout the experiment.

**Apparatus :**

Burette with stand  
Pipette  
1-L volumetric flask with stopper  
250 mL Erlenmeyer flasks  
Retort stand with clamp

**Reagents :**

Sodium hydroxide pellets  
Potassium hydrogen phthalate (KHP)  
Phenolphthalein  
Unknown vinegar

# Sample Introduction Lab Report

**Marco Cascella**



## **Sample Introduction Lab Report:**

## Enjoying the Beat of Expression: An Emotional Symphony within **Sample Introduction Lab Report**

In some sort of eaten by monitors and the ceaseless chatter of quick communication, the melodic splendor and psychological symphony produced by the published word often disappear into the background, eclipsed by the persistent sound and distractions that permeate our lives. However, located within the pages of **Sample Introduction Lab Report** an enchanting fictional value filled with natural feelings, lies an immersive symphony waiting to be embraced. Crafted by a wonderful composer of language, that fascinating masterpiece conducts viewers on an emotional journey, skillfully unraveling the concealed melodies and profound influence resonating within each cautiously constructed phrase. Within the depths with this touching assessment, we shall discover the book is main harmonies, analyze its enthralling writing style, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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