

Average Atomic Mass

How can the masses of the particles tell us something?

Why?

Many different isotopes exist. They are naturally occurring species. As you learned previously, the masses of these isotopes have different values. However, regardless of particular isotopic abundances in the natural element, the mass from different isotopes contributes equal amounts of particles and therefore equal mass. Hence, the mass of the whole element is just the average value. The value obtained is the **average atomic mass**. The average mass of an element is the weighted average of the masses of its isotopes.

Model 1 – A Strip of Magnesium Metal



1. What is the chemical identity of the metal Mg shown in Model 1?
 (How many atoms does it contain?)
2. What are the mass numbers of the isotopes comprising the strip of magnesium shown in Model 1?
 (How many?)
3. How all of the masses of magnesium in Model 1 have the same positive value? Explain.
 (Are there any mass numbers of the atoms that differ? How many atoms are left?)
4. How the sample of 12 isotopes of magnesium shown in Model 1 shows a slight fluctuation in the number of Mg-24 isotopes and the number of atoms of each isotope present.

	Mg-24	Mg-26	Mg-26
Number of atoms	10	2	2

5. Calculate the mass of magnesium in the metal element in Model 1.
 (How?)
 6. Model 1 in Model 1 shows the strip was composed by Element A. The sample of 12 atoms of magnesium represented here shows statistical fluctuations. (How many?)
- As a group of 10 magnesium atoms, there will be approximately one Mg-24 and two Mg-26 atoms.

Periodicity Of Elements Pogil Key

ML Yell



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