

* Notice - the group #1 example given for question #4 contains exceptions to the general trend. Know general trends, not exceptions.

Key P.T.

Periodic Trends

Can the properties of an element be predicted using a periodic table?

Why?

The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

1. Consider the data in Model 1 on the following page.

a. Each element has three numbers listed under it. Which value represents the atomic radius?

The first number under the Electron Shell Diagram

b. What are the units for the atomic radius?

pm - picometer

c. Write a complete sentence to convey your understanding of atomic radius. Note: You may not use the word "radius" in your definition.

The distance between the nucleus and the valence electrons of the atom.

2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.

Atomic radius decreases as you go down a group (vertical column).
Grp #1 Li-152, Na-186, K-227 Grp #2 Ne-70, Ar-98, Kr-112
Grp #3 Be-112, Mg-160, Ca-197

3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. *Hint:* You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.

* *As you go down a group, the greater the distance between the nucleus and valence electrons, the smaller the attractive force.*

4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.

The general trend is atomic radius decreases in a period (horizontal row) moving left to right.

Grp #2 Ne-70, Mg-160, Al-143, Si-117, P-115, S-104, Cl-99, Ar-98

5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.

As you move to the right across a period, protons are added to the nucleus. The greater number of protons increases the attractive force on the valence electrons. This decreases the atomic radius.

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F Rizvi



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Table of Contents Periodic Trends Pogil Worksheet Extension Answers

1. Understanding the eBook Periodic Trends Pogil Worksheet Extension Answers
 - The Rise of Digital Reading Periodic Trends Pogil Worksheet Extension Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Pogil Worksheet Extension Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Pogil Worksheet Extension Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Pogil Worksheet Extension Answers

- Personalized Recommendations
 - Periodic Trends Pogil Worksheet Extension Answers User Reviews and Ratings
 - Periodic Trends Pogil Worksheet Extension Answers and Bestseller Lists
5. Accessing Periodic Trends Pogil Worksheet Extension Answers Free and Paid eBooks
 - Periodic Trends Pogil Worksheet Extension Answers Public Domain eBooks
 - Periodic Trends Pogil Worksheet Extension Answers eBook Subscription Services
 - Periodic Trends Pogil Worksheet Extension Answers Budget-Friendly Options
 6. Navigating Periodic Trends Pogil Worksheet Extension Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Pogil Worksheet Extension Answers Compatibility with Devices
 - Periodic Trends Pogil Worksheet Extension Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Pogil Worksheet Extension Answers
 - Highlighting and Note-Taking Periodic Trends Pogil Worksheet Extension Answers
 - Interactive Elements Periodic Trends Pogil Worksheet Extension Answers
 8. Staying Engaged with Periodic Trends Pogil Worksheet Extension Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Pogil Worksheet Extension Answers
 9. Balancing eBooks and Physical Books Periodic Trends Pogil Worksheet Extension Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Pogil Worksheet Extension Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Trends Pogil Worksheet Extension Answers
 - Setting Reading Goals Periodic Trends Pogil Worksheet Extension Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Trends Pogil Worksheet Extension Answers

- Fact-Checking eBook Content of Periodic Trends Pogil Worksheet Extension Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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