

Periodic Table Basics

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

Step 2: Determine the number of protons, neutrons, and electrons in each element.

Step 3: Create a Bohr diagram for each element.

Step 4: Draw the Lewis Structure for each element.

The card for Boron (B) contains the following information:

- Atomic number: 5
- Symbol: B
- Name: Boron
- Atomic mass: 10.81
- Protons (P): 5
- Neutrons (N): 6
- Electrons (E): 5
- Bohr Diagram: A diagram showing a nucleus with 5 protons and 6 neutrons, and 5 electrons in two shells (2 in the first, 3 in the second).
- Lewis structure: A diagram showing the Boron atom with 5 valence electrons (3 single bonds and 2 lone pairs).

Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

Green = Li & Na

Pink = O & S

Blue = Be & Mg

Purple = F & Cl

Orange = B & Al

Red = C & Si

Tan = N & P

Yellow = He, Ne, & Ar

Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.



Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.

Periodic Table Basics Answer Key

Daniela Niemeyer



Periodic Table Basics Answer Key:

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Table of Contents Periodic Table Basics Answer Key

1. Understanding the eBook Periodic Table Basics Answer Key
 - The Rise of Digital Reading Periodic Table Basics Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Basics Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Basics Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Basics Answer Key
 - Personalized Recommendations
 - Periodic Table Basics Answer Key User Reviews and Ratings
 - Periodic Table Basics Answer Key and Bestseller Lists
5. Accessing Periodic Table Basics Answer Key Free and Paid eBooks
 - Periodic Table Basics Answer Key Public Domain eBooks
 - Periodic Table Basics Answer Key eBook Subscription Services
 - Periodic Table Basics Answer Key Budget-Friendly Options

6. Navigating Periodic Table Basics Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Basics Answer Key Compatibility with Devices
 - Periodic Table Basics Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Basics Answer Key
 - Highlighting and Note-Taking Periodic Table Basics Answer Key
 - Interactive Elements Periodic Table Basics Answer Key
8. Staying Engaged with Periodic Table Basics Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Basics Answer Key
9. Balancing eBooks and Physical Books Periodic Table Basics Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Basics Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Basics Answer Key
 - Setting Reading Goals Periodic Table Basics Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Basics Answer Key
 - Fact-Checking eBook Content of Periodic Table Basics Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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