

CHAPTER 5 REVIEW*The Periodic Law***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- c In the modern periodic table, elements are ordered

 - according to decreasing atomic mass.
 - according to Mendeleev's original design.
 - according to increasing atomic number.
 - based on when they were discovered.
- d Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing

 - density.
 - reactivity.
 - atomic number.
 - atomic mass.
- b The modern periodic law states that

 - no two electrons with the same spin can be found in the same place in an atom.
 - the physical and chemical properties of an element are functions of its atomic number.
 - electrons exhibit properties of both particles and waves.
 - the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
- c The discovery of the noble gases changed Mendeleev's periodic table by adding a new

 - period.
 - series.
 - group.
 - level.
- d The most distinctive property of the noble gases is that they are

 - metallic.
 - radioactive.
 - metalloid.
 - largely unreactive.
- c Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of

 - 4.
 - 10.
 - 11.
 - 18.
- An isotope of fluorine has a mass number of 19 and an atomic number of 9.

 - 9 How many protons are in this atom?
 - 10 How many neutrons are in this atom?
 - ${}^{19}_{9}\text{F}$ What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

Periodic Law Section Quiz Answer Key

Bernard J. Bieg



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Table of Contents Periodic Law Section Quiz Answer Key

1. Understanding the eBook Periodic Law Section Quiz Answer Key
 - The Rise of Digital Reading Periodic Law Section Quiz Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Section Quiz Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Law Section Quiz Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Section Quiz Answer Key
 - Personalized Recommendations
 - Periodic Law Section Quiz Answer Key User Reviews and Ratings
 - Periodic Law Section Quiz Answer Key and Bestseller Lists
5. Accessing Periodic Law Section Quiz Answer Key Free and Paid eBooks

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- Adjustable Fonts and Text Sizes of Periodic Law Section Quiz Answer Key
 - Highlighting and Note-Taking Periodic Law Section Quiz Answer Key
 - Interactive Elements Periodic Law Section Quiz Answer Key
8. Staying Engaged with Periodic Law Section Quiz Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Section Quiz Answer Key
9. Balancing eBooks and Physical Books Periodic Law Section Quiz Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Section Quiz Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Law Section Quiz Answer Key
- Setting Reading Goals Periodic Law Section Quiz Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Law Section Quiz Answer Key
- Fact-Checking eBook Content of Periodic Law Section Quiz Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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