

CAT Sample Paper 1

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Quantitative Ability

DIRECTIONS for questions 1 to 7: Answer the questions independently of each other.

1. The total cost of 2 pencils, 5 erasers and 7 sharpeners is Rs.30, while 3 pencils and 5 sharpeners cost Rs.15 more than 6 erasers. By what amount (in Rs.) does the cost of 39 erasers and 1 sharpener exceed the cost of 6 pencils?

- (1) 20
- (2) 30
- (3) It does not exceed
- (4) Cannot be determined

2. If the roots of the equation $(x + 1)(x + 9) + 8 = 0$ are a and b , then the roots of the equation $(x + a)(x + b) - 8 = 0$ are

- (1) 1 and 9
- (2) -4 and -6
- (3) 4 and 6
- (4) Cannot be determined

3. What is the remainder when 7^{700} is divided by 100?

- (1) 1
- (2) 61
- (3) 41
- (4) 21

4. Balram, the local shoe shop owner, sells four types of footwear - Slippers (S), Canvas Shoes (C), Leather Shoes (L) and Joggers (J). The following information is known regarding the cost prices and selling prices of these four types of footwear:

- (i) L sells for Rs.500 less than J, which costs Rs.300 more than S, which, in turn, sells for Rs.200 more than L.
- (ii) L costs Rs.300 less than C, which sells for Rs.100 more than S.

Sample Paper Of Cat Exam

Michael Seilmaier

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Table of Contents Sample Paper Of Cat Exam

1. Understanding the eBook Sample Paper Of Cat Exam
 - The Rise of Digital Reading Sample Paper Of Cat Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample Paper Of Cat Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Paper Of Cat Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Paper Of Cat Exam
 - Personalized Recommendations

- Sample Paper Of Cat Exam User Reviews and Ratings
- Sample Paper Of Cat Exam and Bestseller Lists
- 5. Accessing Sample Paper Of Cat Exam Free and Paid eBooks
 - Sample Paper Of Cat Exam Public Domain eBooks
 - Sample Paper Of Cat Exam eBook Subscription Services
 - Sample Paper Of Cat Exam Budget-Friendly Options
- 6. Navigating Sample Paper Of Cat Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Sample Paper Of Cat Exam Compatibility with Devices
 - Sample Paper Of Cat Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Paper Of Cat Exam
 - Highlighting and Note-Taking Sample Paper Of Cat Exam
 - Interactive Elements Sample Paper Of Cat Exam
- 8. Staying Engaged with Sample Paper Of Cat Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample Paper Of Cat Exam
- 9. Balancing eBooks and Physical Books Sample Paper Of Cat Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Paper Of Cat Exam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Sample Paper Of Cat Exam
 - Setting Reading Goals Sample Paper Of Cat Exam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sample Paper Of Cat Exam
 - Fact-Checking eBook Content of Sample Paper Of Cat Exam

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 239.8 \text{ g}$ c determine the percent yield $\frac{260.0 \text{ g}}{239.8 \text{ g}} \times 100 = 108.4\%$

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web 7 $\text{Na}_2\text{CO}_3 + 4 \text{HCl} \rightarrow 2 \text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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