

Chemistry – Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. At 22°C , the air pressure in a car tire is 1293 torr. Convert this pressure to atmospheres (atm).
- 12.76
 - 5.262
 - 1.701
 - 1.293

2. This diagram shows a closed-end mercury (Hg) barometer.



When atmospheric pressure is 1.30 atm, what is the height (h), in millimeters, of the Hg in the barometer?

- 132
 - 585
 - 760
 - 988
3. A chemist compresses a 1.5 L sample of carbon dioxide (CO_2) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the CO_2 gas?
- 0.60
 - 0.94
 - 1.7
 - 2.4
4. A vessel connected to an open-end mercury (Hg) manometer contains nitrogen (N_2) gas. The atmospheric pressure is 0.993 atm. The height difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the N_2 gas?

- 355
- 633

- 755
- 877

5. At 1.00 atm, a sealed weather balloon contains 20.0 L of helium (He) gas at 25.0°C . Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at 35.0°C ?
- 40.0
 - 28.0
 - 20.7
 - 14.3
6. At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at 95.9°C . Which statement explains the student's results at this high elevation?
- The boiling point of water increases due to lower atmospheric pressure.
 - The boiling point of water increases due to higher atmospheric pressure.
 - The boiling point of water decreases due to lower atmospheric pressure.
 - The boiling point of water decreases due to higher atmospheric pressure.
7. For a chemistry homework assignment, Beth must consider the effusion rates of methane (CH_4) gas and nitrogen trifluoride (NF_3) gas at 25°C . Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of CH_4 and NF_3 at 25°C ?
- CH_4 effuses 2 times faster than NF_3 .
 - CH_4 effuses 4 times faster than NF_3 .
 - NF_3 effuses 2 times faster than CH_4 .
 - NF_3 effuses 4 times faster than CH_4 .
8. At standard temperature and pressure, a chemist burns 3.50 L of pentane (C_5H_{12}) in excess oxygen (O_2) gas.
- $$\text{C}_5\text{H}_{12}(l) + 8 \text{O}_2(g) \rightarrow 5 \text{CO}_2(g) + 6 \text{H}_2\text{O}(l)$$
- What is the maximum number of liters of carbon dioxide (CO_2) that the reaction can produce?
- 17.5
 - 8.50
 - 5.60
 - 0.700

Sample Test Questions On Gas Laws

**Peter J. Mikulecky, Michelle Rose
Gilman, Kate Brutlag**



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