

Chemistry – Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- At 22°C , the air pressure in a car tire is 1293 torr. Convert this pressure to atmospheres (atm).
a. 12.76
b. 5.262
c. 1.701
d. 1.293

- This diagram shows a closed-end mercury (Hg) barometer.



barometer.

When atmospheric pressure is 1.30 atm, what is the height (h), in millimeters, of the Hg in the barometer?

- 132
 - 585
 - 760
 - 988
- A chemist compresses a 1.5 L sample of carbon dioxide (CO_2) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the CO_2 gas?
a. 0.60
b. 0.94
c. 1.7
d. 2.4
 - A vessel connected to an open-end mercury (Hg) manometer contains nitrogen (N_2) gas. The atmospheric pressure is 0.993 atm. The height difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the N_2 gas?

- 355
- 633

- 755
 - 877
- At 1.00 atm, a sealed weather balloon contains 20.0 L of helium (He) gas at 25.0°C . Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at 35.0°C ?
a. 40.0
b. 28.0
c. 20.7
d. 14.3
 - At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at 95.9°C . Which statement explains the student's results at this high elevation?
a. The boiling point of water increases due to lower atmospheric pressure.
b. The boiling point of water increases due to higher atmospheric pressure.
c. The boiling point of water decreases due to lower atmospheric pressure.
d. The boiling point of water decreases due to higher atmospheric pressure.
 - For a chemistry homework assignment, Beth must consider the effusion rates of methane (CH_4) gas and nitrogen trifluoride (NF_3) gas at 25°C . Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of CH_4 and NF_3 at 25°C ?
a. CH_4 effuses 2 times faster than NF_3 .
b. CH_4 effuses 4 times faster than NF_3 .
c. NF_3 effuses 2 times faster than CH_4 .
d. NF_3 effuses 4 times faster than CH_4 .
 - At standard temperature and pressure, a chemist burns 3.50 L of pentane (C_5H_{12}) in excess oxygen (O_2) gas.
$$\text{C}_5\text{H}_{12}(l) + 8 \text{O}_2(g) \rightarrow 5 \text{CO}_2(g) + 6 \text{H}_2\text{O}(l)$$

What is the maximum number of liters of carbon dioxide (CO_2) that the reaction can produce?
a. 17.5
b. 8.50
c. 5.60
d. 0.700

Sample Test Questions On Gas Laws

Camilla Rothe



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