

Name: _____

Introduction to Solubility PhET Lab (created 3/2011)

We can express concentrations in terms of moles solute per liters of solution

Introduction

When we cook pasta, we frequently dissolve table salt in water. Is there a limit to how much salt we can get into the water?

When ionic-bonded compounds (like table salt, NaCl) are dissolved in polar solvents (like water), the **polar** water molecules act to pull the ions in the compound apart.

When this happens (called dissociation), ions are released into the solution. Is there a limit to the number of ions that can exist in solution?



Salts & Solubility

Pre-lab Vocabulary

Please use your textbook, your notes, or the internet to define the following terms:

Saturated	_____
Unsaturated	_____
Supersaturated	_____
Molarity	_____
Cation	_____
Anion	_____
Dissociation	_____
Solubility	_____

Procedure *PhET Simulations → Play With Sims → Chemistry → Salts and Solubility → Run NaCl*

Table Salt

Take some time and familiarize yourself with the simulation before answering the below questions:

1. What happens to the bound Na⁺ and Cl⁻ ions as they enter the water?

2. Is there a point at which no more sodium and chloride ions will dissolve? How do you know?

3. What do we call a solution that has as many ions as it can hold?

4. What happens if you add additional, solid NaCl after the maximum has been reached?

5. What happens when more water is added to a *saturated* solution?

6. What happens to the ions bound together as a solid when the solution's volume is reduced?

7. How many total Na⁺ ions can you add until no more will dissolve (the maximum amount)?

8. How many **moles** does this many ions represent? (recall, 1 mole = 6.02×10^{23} things)

9. What **volume** (in **Liters**) did you use to get the saturation point in #7 above?

10. What **concentration** is this? (In **mole/L**, use #8 and #9's answers) **This is molarity, M.**



Phet Into To Solubility Simulations Answer Key

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