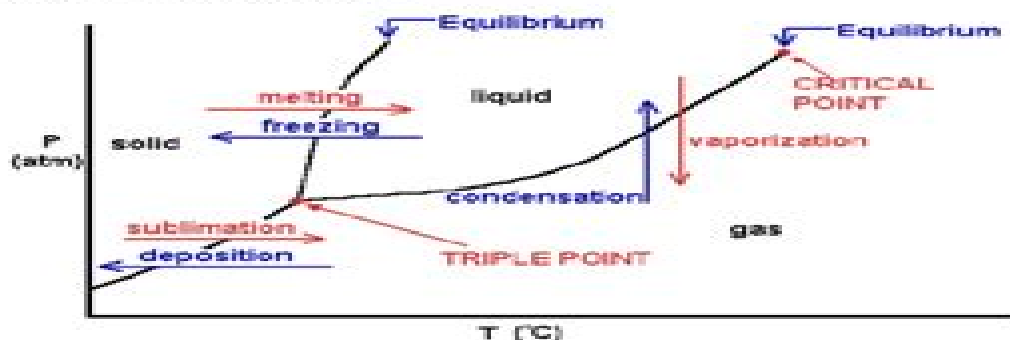


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Chemistry - Changes of State, Vapor Pressure, & Phase Diagrams

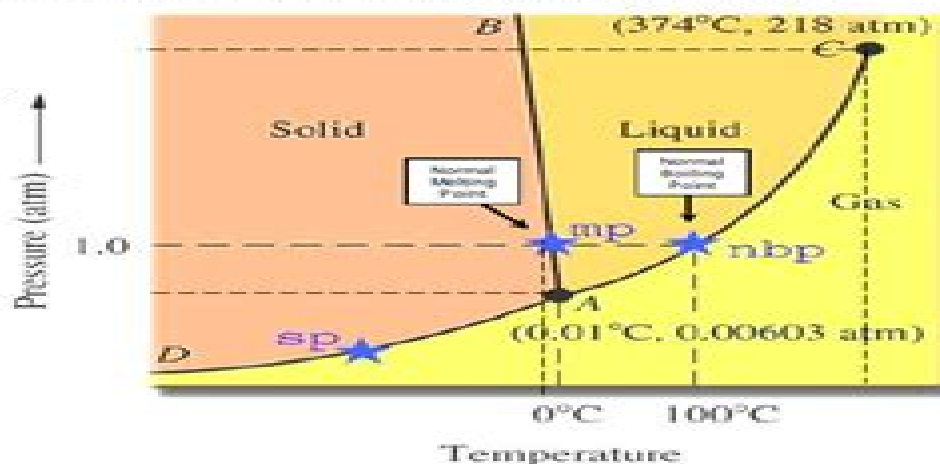
Generic Phase Diagram

- Sublimation is the phase change as a substance changes from a solid to a gas without passing through the intermediate state of a liquid.
- Deposition is the phase change as a substance changes from a gas to a solid without passing through the intermediate state of a liquid.
- **TRIPLE POINT** - The temperature and pressure at which the solid, liquid, and gas phases exist simultaneously.
- **CRITICAL POINT** - The temperature above which a substance will always be a gas regardless of the pressure.
- **NOTE:**
 - The solid phase is more dense than the liquid phase.
 - The line between the solid and gas phases is the equilibrium of solid and gas phases at that specific pressure and temperature, i.e. a curve of all the deposition/sublimation points.
 - The line between the solid and liquid phases is the equilibrium of solid and liquid phases at that specific pressure and temperature, i.e. a curve of all the freezing/melting points.
 - The line between the liquid and gas phases is the equilibrium of liquid and gas phases at that specific pressure and temperature, i.e. a curve of all the vaporization/condensation points.
- **Melting Point (Freezing Point)** - The temperature at which the solid and liquid phases of a substance are in equilibrium at atmospheric pressure.
 - **Normal Melting Point (Freezing Point)** - The temperature at which the solid changes to a liquid at Standard Pressure (1.00 atm = 760 mmHg = 760 torr = 101.325 kPa)
- **Boiling Point (Condensation Point)** - The temperature at which the vapor pressure of a liquid is equal to the pressure on the liquid.
 - **Normal Boiling Point (Condensation Point)** - The temperature at which the vapor pressure of a liquid is equal to Standard Pressure (1.00 atm = 760 mmHg = 760 torr = 101.325 kPa)



Phase Diagram for Water

For water, the liquid phase is more dense than the solid phase due to hydrogen bonding.



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S Baum



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